Stoichiometry Worksheet # 1

 Glucose is used as a source of energy by the human body. The overall reaction in the body is C6H12O6+6O2 → 6CO2 + 6H2O

Calculate the number of moles of oxygen needed to oxidize 12.5 g of C6H12O6.

Ammonia is synthesized from hydrogen and nitrogen according to the following equation.

$$N_2 + 3H_2 \rightarrow 2NH_3$$

If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH₃) can be produced?

3) Assume that in the decomposition of potassium chlorate, KCLO₃,

80.5 moles of O2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,

9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" PtCl₂(NH₂)₂, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction

$$K_2PtCl_4 + 2NH_3 \longrightarrow 2KCl + PtCl_2(NH_3)_2$$

- A. How many moles of "cisplatin" can be produced from 2.50 g K2PtCl4?
- B. How many moles of NH3 would be needed if 3.8 moles of K2PtCl4 were present?

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JL Elias

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Whether systems biology will survive as a field or be washed away by the tides of future fads will ultimately depend on its success to achieve this type of synthesis The famous quote attributed to Kurt Lewin comes to mind there is nothing more practical than a good theory This book presents a wide assortment of articles on systems biology in an attempt to capture the variety of current methods in systems biology and show how they can help to find answers to the challenges of modern biology

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