

Stoichiometry Review

****YOU MUST SHOW ALL WORK AND YOUR ANSWERS MUST INCLUDE THE PROPER NUMBER OF SIG FIGS AND COMPLETE UNITS IN ORDER TO RECEIVE CREDIT FOR THE PROBLEM.****

BALANCE THE FOLLOWING EQUATIONS TO USE IN QUESTIONS 5 – 14:



PERFORM THE FOLLOWING STOICHIOMETRIC CALCULATIONS:

5. Zinc reacts with hydrochloric acid to produce zinc chloride and hydrogen. How many moles of HCl are required to produce 7.50 moles of ZnCl₂?

$$\frac{7.50 \text{ mol ZnCl}_2}{1 \text{ mol ZnCl}_2} \times \frac{2 \text{ mol HCl}}{1 \text{ mol ZnCl}_2} = 15.0 \text{ mol HCl}$$

6. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many grams of copper are required to form 250 g of silver?

$$\frac{250 \text{ g Ag}}{107.868 \text{ g Ag}} \times \frac{1 \text{ mol Ag}}{1 \text{ mol Ag}} \times \frac{1 \text{ mol Cu}}{2 \text{ mol Ag}} \times \frac{63.546 \text{ g Cu}}{1 \text{ mol Cu}} = 74 \text{ g Cu}$$

7. When aluminum is burned in excess oxygen, aluminum oxide is produced. How many grams of oxygen are required to produce 0.75 moles of Al₂O₃?

$$\frac{0.75 \text{ mol Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} \times \frac{3 \text{ mol O}_2}{2 \text{ mol Al}_2\text{O}_3} \times \frac{31.998 \text{ g O}_2}{1 \text{ mol O}_2} = 36 \text{ g O}_2$$

8. How many grams of iron(III) chloride are produced when 15.3 g of iron react with excess chlorine gas?

$$\frac{15.3 \text{ g Fe}}{55.845 \text{ g Fe}} \times \frac{1 \text{ mol Fe}}{1 \text{ mol Fe}} \times \frac{2 \text{ mol FeCl}_3}{2 \text{ mol Fe}} \times \frac{162.204 \text{ g FeCl}_3}{1 \text{ mol FeCl}_3} = 44.4 \text{ g FeCl}_3$$

9. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many moles of silver will be produced from 3.65 moles of silver nitrate?

$$\frac{3.65 \text{ mol AgNO}_3}{2 \text{ mol AgNO}_3} \times \frac{2 \text{ mol Ag}}{1 \text{ mol AgNO}_3} = 3.65 \text{ mol Ag}$$

Stoichiometry Section Review Answer Key

Kaplan Test Prep



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Table of Contents Stoichiometry Section Review Answer Key

1. Understanding the eBook Stoichiometry Section Review Answer Key
 - The Rise of Digital Reading Stoichiometry Section Review Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Section Review Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Section Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Section Review Answer Key

- Personalized Recommendations
 - Stoichiometry Section Review Answer Key User Reviews and Ratings
 - Stoichiometry Section Review Answer Key and Bestseller Lists
5. Accessing Stoichiometry Section Review Answer Key Free and Paid eBooks
 - Stoichiometry Section Review Answer Key Public Domain eBooks
 - Stoichiometry Section Review Answer Key eBook Subscription Services
 - Stoichiometry Section Review Answer Key Budget-Friendly Options
 6. Navigating Stoichiometry Section Review Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Section Review Answer Key Compatibility with Devices
 - Stoichiometry Section Review Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Section Review Answer Key
 - Highlighting and Note-Taking Stoichiometry Section Review Answer Key
 - Interactive Elements Stoichiometry Section Review Answer Key
 8. Staying Engaged with Stoichiometry Section Review Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Section Review Answer Key
 9. Balancing eBooks and Physical Books Stoichiometry Section Review Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Section Review Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Stoichiometry Section Review Answer Key
 - Setting Reading Goals Stoichiometry Section Review Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Stoichiometry Section Review Answer Key

- Fact-Checking eBook Content of Stoichiometry Section Review Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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