

Working a Stoichiometry Problem

6.50 grams of aluminum reacts with an excess of oxygen. How many grams of aluminum oxide are formed?



$$6.50 \times 2 \times 101.96 \div 26.98 \div 4 = 12.3 \text{ g Al}_2\text{O}_3$$

6.50 g Al	1 mol Al	2 mol Al₂O₃	101.96 g Al₂O₃
	26.98 g Al	4 mol Al	1 mol Al₂O₃

$$= ? \text{ g Al}_2\text{O}_3$$

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