

Stoichiometry Worksheet and Key



1. How many moles of O_2 will be formed from 1.65 moles of KClO_3 ?

1.65 mol KClO_3	_____ mol O_2	= _____ mol O_2
	_____ mol KClO_3	

2. How many moles of KClO_3 are needed to make 3.50 moles of KCl ?

3.50 mol KCl		= _____ mol KClO_3

3. How many moles of KCl will be formed from 2.73 moles of KClO_3 ?

_____		= _____



4. How many moles of Fe_2O_3 are produced when 0.275 moles of Fe is reacted?

0.275 mol Fe		= _____ mol Fe_2O_3

5. How many moles of Fe_2O_3 are produced when 31.0 moles of O_2 is reacted?

_____		= _____

6. How many moles of O_2 are needed to react with 8.9 moles of Fe ?

_____		= _____



7. How many moles of O_2 are produced when 1.26 moles of H_2O is reacted?

8. How many moles of H_2O are needed to produce 55.7 moles of H_2 ?

9. If enough H_2O is reacted to produce 3.40 moles of H_2 , then how many moles of O_2 must have been made?
(a bit challenging, but just think about it and you can probably figure it out)

Stoichiometry Problem Sheet 1

Siddappa N.Byraredy



Stoichiometry Problem Sheet 1:

Elements of Environmental Chemistry Ronald A. Hites, Jonathan D. Raff, 2013-08-26 From Reviews of the First Edition This splendid at times humorous and reasonably priced little book has much to commend it to undergraduate chemists and to other science students J G Farmer University of Edinburgh Complex environmental issues are presented in simple terms to help readers grasp the basics and solve relevant problems J Albaiges University of Barcelona The main strength of the book lies in its explanations of the calculation of quantitative relationships Each chapter includes 15 20 problems that are carefully chosen from a didactic standpoint for which the reader can find solutions at the end D Lenoir Institute for Ecological Chemistry What drew me to the first edition was the style the no nonsense down to earth explanations and the practical examples that litter the text The dry humor expressed in the footnotes is great and reminds me of other classic texts T Clough Lincoln University A practical approach to environmental chemistry Providing readers with the fundamentals of environmental chemistry and a toolbox for putting them into practice **Elements of Environmental Chemistry Second Edition** is a concise accessible and hands on volume designed for students and professionals working in the chemical and environmental sciences Tutorial in style this book fully incorporates real world problems and extensive end of chapter problem sets to immerse the reader in the field Chapters cover mass balance chemical kinetics carbon dioxide equilibria pesticide structures and much more Extensively revised updated and expanded this Second Edition includes new chapters on atmospheric chemistry climate change and polychlorinated biphenyls and dioxins and brominated flame retardants In addition new practice problems and a helpful tutorial on organic chemistry names and structures have been added to improve both the scope and accessibility of the book

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undergraduate and graduate students studying mechanical electrical and electronic engineering *Proceedings*, 1978

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