

Name: Key

Class: _____

Date: _____

Stoichiometry Problem Set A

Directions: Solve each of the following problems. Show work using the factor label method or the irf box. Include correct units and significant figures with each answer. Clearly indicate final answer. Be sure to start with a balanced equation

Part 1: Straight Stoichiometry Problems

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method.

- 1.) For the reaction $2\text{HNO}_3 + 1\text{Mg(OH)}_2 \rightarrow 1\text{Mg(NO}_3)_2 + 2\text{H}_2\text{O}$, how many grams of magnesium ^B nitrate are produced from 8.00 mol of nitric acid (HNO_3)? A

$$\text{MM} = 148.3\text{g}$$

$$\frac{8.00\text{mol HNO}_3}{1} \times \frac{1\text{mol Mg(NO}_3)_2}{2\text{mol HNO}_3} \times \frac{148.3\text{g Mg(NO}_3)_2}{1\text{mol Mg(NO}_3)_2} = 592.8\text{g Mg(NO}_3)_2$$

- 2.) For the reaction $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$, how many moles of potassium ^B chlorate are required to produce 160. g of oxygen? A

$$\text{MM} = 32.0\text{g}$$

$$\frac{160.\text{g O}_2}{1} \times \frac{1\text{mol O}_2}{32.0\text{g O}_2} \times \frac{2\text{mol KClO}_3}{3\text{mol O}_2} = 3.33\text{ mole KClO}_3$$

- 3.) For the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, how many grams of water are produced from 5.22 liters of oxygen gas? ^B A

$$\text{MM} = 18.0\text{g}$$

$$\frac{5.22\text{L O}_2}{1} \times \frac{1\text{mol O}_2}{22.4\text{L O}_2} \times \frac{2\text{mol H}_2\text{O}}{1\text{mol O}_2} \times \frac{18.0\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 8.39\text{g H}_2\text{O}$$

- 4.) In the reaction $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$, how many grams of hydrogen ^B are produced if 1.33 grams of sodium react completely? A

$$\text{MM} = 23.0\text{g}$$

$$\text{MM} = 2.0\text{g}$$

$$\frac{1.33\text{g Na}}{1} \times \frac{1\text{mol Na}}{23.0\text{g Na}} \times \frac{1\text{mole H}_2}{2\text{mole Na}} \times \frac{2.0\text{g H}_2}{1\text{mole H}_2} = 0.0578\text{g H}_2$$

- 5.) In the reaction $1\text{CO}_2 + 1\text{H}_2\text{O} \rightarrow 1\text{H}_2\text{CO}_3$, how many moles of carbonic ^B acid can be produced from 97.2 g of CO_2 ? A

$$\text{MM} = 44.0\text{g}$$

$$\frac{97.2\text{g CO}_2}{1} \times \frac{1\text{mole CO}_2}{44.0\text{g CO}_2} \times \frac{1\text{mole H}_2\text{CO}_3}{1\text{mole CO}_2} = 2.21\text{ moles H}_2\text{CO}_3$$

Did you achieve our target goal?

Target Goal: To properly calculate the amount of one substance equal to a given amount of another substance within a chemical reaction using the factor-label method. Circle your answer.

1
Feeling VERY confused

2
Not sure that I get this

3
I think I am catching on

4
Doing great! Bring it on!

Stoichiometry Problem 1 Answers

**Robert J. Salm, Robert J. Salm Robert J.
Salm Ph.D.**

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Microbiology (Questions and Answers), 5e Purshotam Kaushik & Kirti Kaushik, 2022 Microbiology is an engaging textbook presenting balanced and comprehensive account of major areas of microbiology in the form of questions and answers This question answer approach to present complex topics and theories of microbiology regarding cellular and non cellular microorganisms microbial genetics and molecular biology in higher plants and animals makes the subject interesting and easily comprehensible for the students Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests *Microbiology Questions and Answers* Mr. Ajeet Singh, 2018-05-14 A structured question and answer resource covering core microbiology concepts laboratory techniques microbial physiology genetics and applied microbiology for academic and competitive exams *Jacaranda Chemistry 1 VCE Units 1 and 2, learnON and Print* Neale Taylor, Angela Stubbs, Robert Stokes, 2022-11-30 Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023 2027

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water and the hydrosphere which also covers water pollution and water purification Water purification occurs in nature through the usual water cycle and by the action of microorganisms The air flushes dissolved gases and volatile pollutants when water seeps through the soil it filters solids as they settle in the bottom of placid lakes Microorganisms break down large organic molecules containing mostly carbon hydrogen nitrogen oxygen sulfur or phosphorus into harmless molecules and ions This text notes that natural purification occurs if the level of contaminants is not so excessive This textbook is suitable for both chemistry teachers and students Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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