

Stoichiometry Practice Worksheet

Balancing Equations and Simple Stoichiometry

Balance the following equations:

- 1) $\text{___ N}_2 + \text{___ F}_2 \rightarrow \text{___ NF}_3$
- ✓ 2) $\text{___ C}_6\text{H}_{10} + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- ✓ 3) $\text{___ HBr} + \text{___ KHCO}_3 \rightarrow \text{___ H}_2\text{O} + \text{___ KBr} + \text{___ CO}_2$
- ✓ 4) $\text{___ GaBr}_3 + \text{___ Na}_2\text{SO}_3 \rightarrow \text{___ Ga}_2(\text{SO}_3)_3 + \text{___ NaBr}$
- 5) $\text{___ SnO} + \text{___ NF}_3 \rightarrow \text{___ SnF}_2 + \text{___ N}_2\text{O}_3$

Solve the following stoichiometry grams-grams problems:

- ✓ 6) Using the following equation:



How many grams of sodium sulfate will be formed if you start with 200 grams of sodium hydroxide and you have an excess of sulfuric acid?

↗ Using the following equation:



How many grams of lithium nitrate will be needed to make 250 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction?

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Sassy Stoichiometry Problems Julie C. Gilbert, 2021-03-14 Need more Stoichiometry practice Stoichiometry has been striking fear into the hearts of chemistry students for ages The best way to conquer something is to practice it Inside you'll find Brief descriptions of each type of ideal stoichiometry and limiting reactant stoichiometry 4 ideal stoichiometry worksheets broken down by type with keys and explanations 4 ideal stoichiometry self quizzes with their answer keys 2 limiting reactant stoichiometry worksheets with keys and explanations 2 limiting reactant stoichiometry self quizzes with answer keys 2 mixed stoichiometry self tests with answer keys This is a companion workbook for the 5 Steps to Surviving Chemistry book However you do not need to have read that book to find this workbook useful **A Stoichiometry Unit** David Callaghan, 2004 *Supporting Traditional Methods with a Constructivist Approach to Learning* Kenneth Munoz Abayan, ProQuest Dissertations and Theses (Electronic resource collection), 2014 Stoichiometry is a fundamental topic in chemistry that measures a quantifiable relationship between atoms molecules etc Stoichiometry is usually taught using expository teaching methods Students are passively given information in the hopes they will retain the transmission of information to be able to solve stoichiometry problems masterfully Cognitive science research has shown that this kind of instructional teaching method is not very effective in meaningful learning practice Instead students must take ownership of their learning The students need to actively construct their own knowledge by receiving interpreting integrating and reorganizing that information into their own mental schemas In the absence of active learning practices tools must be created in such a way to be able to scaffold difficult problems by encoding opportunities necessary to make the construction of knowledge memorable thereby creating a usable knowledge base Using an online e learning tool and its potential to create a dynamic and interactive learning environment may facilitate the learning of stoichiometry The study entailed requests from volunteer students IRB consent form a baseline questionnaire random assignment of treatment pre and post test assessment and post assessment survey These activities were given online A stoichiometry based assessment was given in a proctored examination at the University of Texas at Arlington UTA campus The volunteer students who took part in these studies were at least 18 of age and were enrolled in General Chemistry 1441 at the University of Texas at Arlington Each participant gave their informed consent to use their data in the following study Students were randomly assigned to one of 4 treatment groups based on teaching methodology Dimensional Analysis Operational Method Ratios and Proportions and a control group who just received instruction through lecture only In this study an e learning tool was created to demonstrate several methodologies on how to solve stoichiometry which are all supported by chemical education research Comparisons of student performance based on pre and post test assessment and a stoichiometry based examination was done to determine if the information provided within the e learning tool yielded greater learning outcomes compared to the students in the absence of scaffold learning material The e learning tool was created to help scaffold the problem solving process necessary to help

students N 394 solve stoichiometry problems. Therein the study investigated possible predictors for success on a stoichiometry based examination: students' conceptual understanding of solving stoichiometry problems and their explanation of reasoning. It was found that the way the student answered a given stoichiometry question, i.e. whether the student used dimensional analysis, operational method or any other process, was not statistically relevant ($p > 0.05$). More importantly, if the students were able to describe their thought process clearly, these students scored significantly higher on stoichiometry test (mean 84%).

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