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Chemistry: Stoichiometry – Problem Sheet 1

Directions: Solve each of the following problems. Show your work, including proper units, to earn full credit.

1. Silver and nitric acid react according to the following balanced equation:



- A. How many moles of silver are needed to react with 40 moles of nitric acid?
- B. From the amount of nitric acid given in Part A, how many moles of silver nitrate will be produced?
- C. From the amount of nitric acid given in Part A, how many moles of water will be produced?
- D. From the amount of nitric acid given in Part A, how many moles of nitrogen monoxide will be made?

2. Given the balanced equation: $2 \text{ N}_2\text{H}_4\text{(l)} + \text{ N}_2\text{O}_4\text{(l)} \rightarrow 3 \text{ N}_2\text{(g)} + 4 \text{ H}_2\text{O(g)}$

- A. How many moles of dinitrogen tetrahydride are required to produce 57 moles of nitrogen?
- B. How many moles of dinitrogen tetroxide are required to produce 57 moles of nitrogen?
- C. How many moles of water are produced when 57 moles of nitrogen are made?

3. Calculate the mass of aluminum oxide produced when 3.75 moles of aluminum burn in oxygen.

Answers:

1A. 30 mol Ag
1B. 30 mol AgNO₃

1C. 20 mol H₂O
1D. 10 mol NO

2A. 38 mol N₂H₄
2B. 19 mol N₂O₄

2C. 76 mol H₂O
3. 191 g Al₂O₃

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SA Adler



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Table of Contents Stoichiometry Guided Study Work

1. Understanding the eBook Stoichiometry Guided Study Work
 - The Rise of Digital Reading Stoichiometry Guided Study Work
 - Advantages of eBooks Over Traditional Books
2. Identifying Stoichiometry Guided Study Work
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Guided Study Work
 - User-Friendly Interface
4. Exploring eBook Recommendations from Stoichiometry Guided Study Work
 - Personalized Recommendations
 - Stoichiometry Guided Study Work User Reviews and Ratings
 - Stoichiometry Guided Study Work and Bestseller Lists

5. Accessing Stoichiometry Guided Study Work Free and Paid eBooks
 - Stoichiometry Guided Study Work Public Domain eBooks
 - Stoichiometry Guided Study Work eBook Subscription Services
 - Stoichiometry Guided Study Work Budget-Friendly Options
6. Navigating Stoichiometry Guided Study Work eBook Formats
 - ePub, PDF, MOBI, and More
 - Stoichiometry Guided Study Work Compatibility with Devices
 - Stoichiometry Guided Study Work Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Guided Study Work
 - Highlighting and Note-Taking Stoichiometry Guided Study Work
 - Interactive Elements Stoichiometry Guided Study Work
8. Staying Engaged with Stoichiometry Guided Study Work
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Guided Study Work
9. Balancing eBooks and Physical Books Stoichiometry Guided Study Work
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Guided Study Work
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Stoichiometry Guided Study Work
 - Setting Reading Goals Stoichiometry Guided Study Work
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Stoichiometry Guided Study Work
 - Fact-Checking eBook Content of Stoichiometry Guided Study Work
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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