CHEMISTRY STOCHIOMETRY EXAM LEVEL 1 & 2

1. (a) A student carries out a titration to find the concentration of some sulfuric acid.

The student finds that 25.00 cm³ of 0.0880 mol dm⁻³ aqueous sodium hydroxide, NaOH, is neutralised by 17.60 cm³ of dilute sulfuric acid, H₂SO₄.

 $H_2SO_4(aq) + 2NaOH(aq) \rightarrow Na_2SO_4(aq) + 2H_2O(I)$

(i) Calculate the amount, in moles, of NaOH used.

answer =mol

- [7]
- (ii) Determine the amount, in moles, of H₂SO₄ used.

answer =mol

(iii) Calculate the concentration, in mol dm-3, of the sulfuric acid.

answer =mol dm-3

- (b) After carrying out the titration in (a), the student left the resulting solution to crystallise. White crystals were formed, with a formula of Na₂SO₄•x H₂O and a molar mass of 322.1 g mol⁻¹.
 - (i) What term is given to the '•x H₂O' part of the formula?

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Stoichiometry Exam V

L.R. Morss, Norman M. Edelstein, Jean Fuger

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Nuclear Science and Engineering ,1986

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Table of Contents Stoichiometry Exam V

- 1. Understanding the eBook Stoichiometry Exam V
 - The Rise of Digital Reading Stoichiometry Exam V
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Stoichiometry Exam V
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Stoichiometry Exam V
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Stoichiometry Exam V
 - Personalized Recommendations
 - Stoichiometry Exam V User Reviews and Ratings

- Stoichiometry Exam V and Bestseller Lists
- 5. Accessing Stoichiometry Exam V Free and Paid eBooks
 - Stoichiometry Exam V Public Domain eBooks
 - Stoichiometry Exam V eBook Subscription Services
 - Stoichiometry Exam V Budget-Friendly Options
- 6. Navigating Stoichiometry Exam V eBook Formats
 - o ePub, PDF, MOBI, and More
 - Stoichiometry Exam V Compatibility with Devices
 - Stoichiometry Exam V Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Stoichiometry Exam V
 - Highlighting and Note-Taking Stoichiometry Exam V
 - Interactive Elements Stoichiometry Exam V
- 8. Staying Engaged with Stoichiometry Exam V
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Stoichiometry Exam V
- 9. Balancing eBooks and Physical Books Stoichiometry Exam V
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Stoichiometry Exam V
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Stoichiometry Exam V
 - Setting Reading Goals Stoichiometry Exam V
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Stoichiometry Exam V
 - Fact-Checking eBook Content of Stoichiometry Exam V
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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