2.355a Cu (C2H3Oz)2.	L	mal Cu (Człły	g)2	1000 ML solu
2.355g Cn (C2H3Oz) 2 -	181.2	g Cu(CzH3¢	2)2	1 L sola
= 0.259504/	v = p	.260 M		
Sample Problem 2: Calculate the concentration in ppm. p				
copper (II) acetate (FW 181.5) and dil density of solution at the given tempe 4.70 to 104 ppm; 4.70 x 104 ppth		043 o/ml		flask. The
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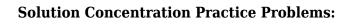
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