

## Solubility Product, $K_{sp}$ - Worksheet 1

1) Write the dissociation equation and the solubility product ( $K_{sp}$ ) expression for the dissociation of each of the following substances.

- a)  $PbSO_4$  (s)  
b)  $Mg(OH)_2$  (s)

- c)  $Pb(NO_3)_2$  (s)  
d)  $AlCl_3$  (s)

- e)  $Na_2S$  (s)  
f)  $Ca_3(PO_4)_2$  (s)

2) Write the solubility product expression for the dissociation of each of the following solids in water.

- a) sodium chloride  
b) calcium carbonate

- c) silver sulfide  
d) aluminum hydroxide

- e) magnesium sulfite  
f) barium phosphate

3) The solubility of  $NiS$  is  $1.0 \times 10^{-12}$  mol/L. Calculate the value of  $K_{sp}$  for this compound.

4) The solubility of calcium hydroxide ( $Ca(OH)_2$ ) is 0.02 mol/L. Calculate the solubility product constant for this substance.

5) The solubility of silver sulfide ( $Ag_2S$ ) is  $1.3 \times 10^{-6}$  mol/L. Calculate the  $K_{sp}$  value.

6) Experiments show that 0.0059 g of  $SrCO_3$  will dissolve in 1.0 L of water. What is the  $K_{sp}$  for strontium carbonate?

7) The solubility of cobalt (II) sulfide is  $3.8 \times 10^{-3}$  g/L. Calculate the value of  $K_{sp}$ .

## Solubility Product, $K_{sp}$ - Worksheet 2

1. Calculate the solubility (in mol/L) for each of the following compounds:

- |               |                                |
|---------------|--------------------------------|
| a. $FeS$      | $K_{sp} = 3.7 \times 10^{-19}$ |
| b. $Mg(OH)_2$ | $K_{sp} = 9.0 \times 10^{-12}$ |
| c. $Ag_2S$    | $K_{sp} = 1.6 \times 10^{-49}$ |
| d. $AgCN$     | $K_{sp} = 2.0 \times 10^{-12}$ |

2. What is the solubility of silver iodide in grams per litre? ( $K_{sp} = 8.3 \times 10^{-17}$ )

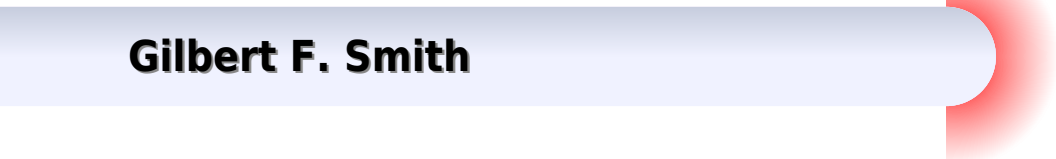
3. The  $K_{sp}$  value for  $BaSO_4$  is  $1.5 \times 10^{-9}$ . Calculate the number of moles of this salt that will dissolve in 1000 mL of water.

4. Calculate the mass of  $CuS$  that will dissolve in 1.0 litre of water.  $K_{sp} = 4.0 \times 10^{-38}$

5. Calculate the mass of barium carbonate that will dissolve in 100 mL of water. The  $K_{sp}$  of  $BaCO_3$  is  $8.1 \times 10^{-9}$ .

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**Gilbert F. Smith**



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## **Table of Contents Solubility Product Answers Chemfiesta**

1. Understanding the eBook Solubility Product Answers Chemfiesta
  - The Rise of Digital Reading Solubility Product Answers Chemfiesta
  - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Product Answers Chemfiesta
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Solubility Product Answers Chemfiesta
  - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Product Answers Chemfiesta
  - Personalized Recommendations
  - Solubility Product Answers Chemfiesta User Reviews and Ratings
  - Solubility Product Answers Chemfiesta and Bestseller Lists
5. Accessing Solubility Product Answers Chemfiesta Free and Paid eBooks
  - Solubility Product Answers Chemfiesta Public Domain eBooks
  - Solubility Product Answers Chemfiesta eBook Subscription Services
  - Solubility Product Answers Chemfiesta Budget-Friendly Options
6. Navigating Solubility Product Answers Chemfiesta eBook Formats

- ePub, PDF, MOBI, and More
- Solubility Product Answers Chemfiesta Compatibility with Devices
- Solubility Product Answers Chemfiesta Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Solubility Product Answers Chemfiesta
  - Highlighting and Note-Taking Solubility Product Answers Chemfiesta
  - Interactive Elements Solubility Product Answers Chemfiesta
- 8. Staying Engaged with Solubility Product Answers Chemfiesta
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Solubility Product Answers Chemfiesta
- 9. Balancing eBooks and Physical Books Solubility Product Answers Chemfiesta
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Solubility Product Answers Chemfiesta
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Solubility Product Answers Chemfiesta
  - Setting Reading Goals Solubility Product Answers Chemfiesta
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solubility Product Answers Chemfiesta
  - Fact-Checking eBook Content of Solubility Product Answers Chemfiesta
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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