RB	#	Name	
Ro	v/Seat		Per
	Sol	ubility Curves	
The	following graph shows solubility curv	ves. Study the graph then answ	er the following questions.
1.	What relationship exists between solu and temperature for most of the sub st shown?		
2a.	What is the exception?	80	
ь	What general principle accounts for this exception?	04 60 (A	\$ 15°
За.	Approximately how many grams of NaNO <sub>3</sub> will dissolve in 100 g water at 10 °C?	198 50 50 40	KCI
b.	How many grams will dissolve at 34 °C?	Aggento 30 NaCl	¥clo.
4.	How many grams of KCl will dissolve in 1 liter of H <sub>2</sub> O at 50 °C?	e o 20	AVI,
5.	58.0 g of K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> is added to 100 g H 0 °C. With constant stirring, to what t erature must the solution be raised to p a saturated solution with no solid K <sub>2</sub> C	temp- 0 10 20 30 40	50 60 70 80 90 100 mperature *C
6.	A saturated solution of KClO <sub>3</sub> was ma 34 °C. How much KClO <sub>3</sub> could be re solution to dryness?	ade with 300 g of H <sub>2</sub> O at covered by evaporating the	answer on left     2a.
7.	500 g of water are used to make a satu 15 °C. How many more grams of KC temperature were raised to 78 °C?		b. answer on left
8	A saturated solution of KNO <sub>3</sub> in 200 c	g of H₂O at 51 °C is cooled	3a
	to 28 °C. How much KNO3 will preci	ipitate out of solution?	b
1.			4
			5.
			6
zь.			7

8.\_\_\_\_

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**RS Peters** 

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