

CHAPTER 1: DATABASES AND DATABASE USERS

Answers to Selected Exercises

1.8 - Identify some informal queries and update operations that you would expect to apply to the database shown in Figure 1.2.

Answer:

(a) (Query) List the names of all students majoring in Computer Science.

(b) (Query) What are the prerequisites of the Database course?

(c) (Query) Retrieve the transcript of Smith. This is a list of <CourseName, SectionIdentifier, Semester, Year, Grade> for each course section that Smith has completed.

(d) (Update) Insert a new student in the database whose Name=Jackson, StudentNumber=23, Class=1 (freshman), and Major=MATH.

(e) (Update) Change the grade that Smith received in Intro to Computer Science section 119 to B.

1.9 - What is the difference between controlled and uncontrolled redundancy?

Answer:

Redundancy is when the same fact is stored multiple times in several places in a database. For example, in Figure 1.5(a) the fact that the name of the student with StudentNumber=8 is Brown is stored multiple times. Redundancy is controlled when the DBMS ensures that multiple copies of the same data are consistent; for example, if a new record with StudentNumber=8 is stored in the database of Figure 1.5(a), the DBMS will ensure that StudentName=Smith in that record. If the DBMS has no control over this, we have uncontrolled redundancy.

1.10 - Specify all the relationships among the records of the database shown in Figure 1.2.

Answer:

(a) Each SECTION record is related to a COURSE record.

(b) Each GRADE_REPORT record is related to one STUDENT record and one SECTION record.

(c) Each PREREQUISITE record relates two COURSE records: one in the role of a course and the other in the role of a prerequisite to that course.

1.11 - Give some additional views that may be needed by other user groups for the database shown in Figure 1.2.

Answer:

(a) A view that groups all the students who took each section and gives each student's grade. This may be useful for printing the grade report for each section for the university administration's use.

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web apr 7 2014 molarity moles of solute liters of solution 8.422 a first convert 250 ml to liters 250/1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3. c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture can be calculated using the following formula $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$ $\frac{480 \text{ g} \times 1.2 \text{ L} + 520 \text{ g} \times 2.0 \text{ L} + 624 \text{ g} \times 1.0 \text{ L}}{1.2 \text{ L} + 2.0 \text{ L} + 1.0 \text{ L}} = 544 \text{ g/L}$ with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1. calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution 2. what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3. 140 g of KCl is dissolved in 600 ml of water what is the molarity 4. 724.4 g of ammonium phosphate in 4500 ml of alcohol

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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