

ASSIGNMENT 3 CE SOLUTION MANUAL

- P18.13** (a) The original volume of the acetone we take as precisely 100 mL. After it is finally cooled to 20.0°C, its volume is

$$V_f = V_i(1 + \beta \Delta T) = (100 \text{ mL}) \left\{ 1 + \left[1.50 \times 10^{-4} (\text{°C})^{-1} \right] (-15.0 \text{°C}) \right\} \\ = \boxed{99.8 \text{ mL}}$$

- (b) Initially, the volume of the acetone reaches the 100-mL mark on the flask, but the acetone cools and the flask warms to a temperature of 32.0 °C. Thus, the volume of the acetone decreases and the volume of the flask increases. This means the acetone will be below the 100-mL mark on the flask.

- P18.40** (a) Particle in equilibrium model

- (b) On the piston,

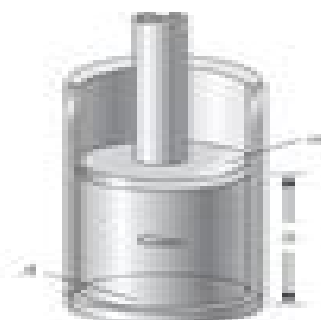
$$\sum F = F_{\text{gas}} - F_g - F_{\text{ext}} = 0 \\ \sum F = PA - mg - P_{\text{atm}}A = 0$$

- (c) In equilibrium, $P_{\text{gas}} = \frac{mg}{A} + P_{\text{atm}}$

$$\text{Therefore, } \frac{nRT}{hA} = \frac{mg}{A} + P_{\text{atm}}$$

$$\text{or } \boxed{h = \frac{nRT}{mg + P_{\text{atm}}A}}$$

where we have used $V = hA$ as the volume of the gas.



ANS. FIG. P18.40

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Ralph H. Petrucci, Robert K. Wismer



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Table of Contents Solution Manual Ece

1. Understanding the eBook Solution Manual Ece
 - The Rise of Digital Reading Solution Manual Ece
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Ece
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Ece
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Ece
 - Personalized Recommendations
 - Solution Manual Ece User Reviews and Ratings
 - Solution Manual Ece and Bestseller Lists
5. Accessing Solution Manual Ece Free and Paid eBooks

- Solution Manual Ece Public Domain eBooks
- Solution Manual Ece eBook Subscription Services
- Solution Manual Ece Budget-Friendly Options
- 6. Navigating Solution Manual Ece eBook Formats
 - ePub, PDF, MOBI, and More
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 - Solution Manual Ece Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Manual Ece
 - Highlighting and Note-Taking Solution Manual Ece
 - Interactive Elements Solution Manual Ece
- 8. Staying Engaged with Solution Manual Ece
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Ece
- 9. Balancing eBooks and Physical Books Solution Manual Ece
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Ece
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Solution Manual Ece
 - Setting Reading Goals Solution Manual Ece
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Solution Manual Ece
 - Fact-Checking eBook Content of Solution Manual Ece
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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