

**Solution Manual for Chemistry Principles and Reactions 8th Edition by
Masterton Hurley ISBN 130507937X 9781305079373**

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2 ATOMS, MOLECULES AND IONS

LECTURE NOTES

Students find this material relatively easy to assimilate; it's almost entirely qualitative. On the other hand, there's a lot of memorizing (sorry, learning) to do. This chapter is coverable in two lectures.

Some general observations:

- ✓ Material in Sections 2.1–2.3 is generally well covered in high-school chemistry courses; no need to dwell on it.
- ✓ Students need to know the molecular formulas of the elements (Figure 2.13), the charges of ions with noble-gas structures and the names and formulas of the common polyatomic ions (Table 2.2). The charges of transition-metal ions will be covered later, in Chapter 4.
- ✓ Naming compounds requires students to distinguish between ionic and molecular substances. It helps to point out that binary molecular compounds are composed of two nonmetals. Almost all ionic compounds contain a metal cation combined with a nonmetal anion or negatively charged polyatomic ion. The flow charts shown in Figures 2.18 and 2.19 should help visual learners.
- ✓ The periodic table will be discussed in greater detail later in the text (Chapter 6).

Lecture 1

I. Atomic Theory

A. Elements

Postulates: Elements consist of tiny particles called atoms, which retain their identity in reactions. In a compound, atoms of two or more elements combine in a fixed ratio of small whole numbers (e.g., 1:1, 2:1, etc.).

B. Components

	relative mass	relative charge	location
proton	1	+1	nucleus
neutron	1	0	nucleus
electron	0.0005	1	outside

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