

Student Solution Manual

Student Solution Manual: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45055-0, Chapter 2: Calculations Used in Analytical Chemistry

Some of the answers below may differ in format but have the same value as your result. Please check with your instructor if a specific format is desired.

Chapter 2

2-1. Define

Answers:

(a) molar mass.

The molar mass is the mass in grams of one mole of a chemical species.

(c) millimolar mass.

The millimolar mass is the mass in grams of one millimole of a chemical species.

2-3. Give two examples of units derived from the fundamental base SI units.

Solution:

$$\text{The liter: } 1 \text{ L} = \frac{1000 \text{ mL}}{1 \text{ L}} \times \frac{1 \text{ cm}^3}{\text{mL}} \times \left(\frac{\text{m}}{100 \text{ cm}} \right)^3 = 10^{-3} \text{ m}^3$$

$$\text{Molar concentration: } 1 \text{ M} = \frac{1 \text{ mol}}{\text{L}} \times \frac{\text{L}}{10^{-3} \text{ m}^3} = \frac{1 \text{ mol}}{10^{-3} \text{ m}^3}$$

2-4. Simplify the following quantities using a unit with an appropriate prefix:

Solutions:

(a) $5.8 \times 10^8 \text{ Hz}$.

$$5.8 \times 10^8 \text{ Hz} \times \frac{\text{MHz}}{10^6 \text{ Hz}} = 580 \text{ MHz}$$

(c) $9.31 \times 10^7 \mu\text{mol}$.

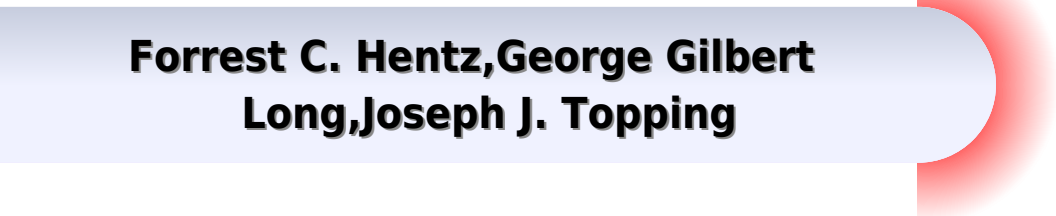
$$9.31 \times 10^7 \mu\text{mol} \times \frac{\text{mol}}{10^6 \mu\text{mol}} = 93.1 \text{ mol}$$

(e) $3.96 \times 10^6 \text{ nm}$.

$$3.96 \times 10^6 \text{ nm} \times \frac{\text{mm}}{10^6 \text{ nm}} = 3.96 \text{ mm}$$

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**Forrest C. Hentz, George Gilbert
Long, Joseph J. Topping**



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