

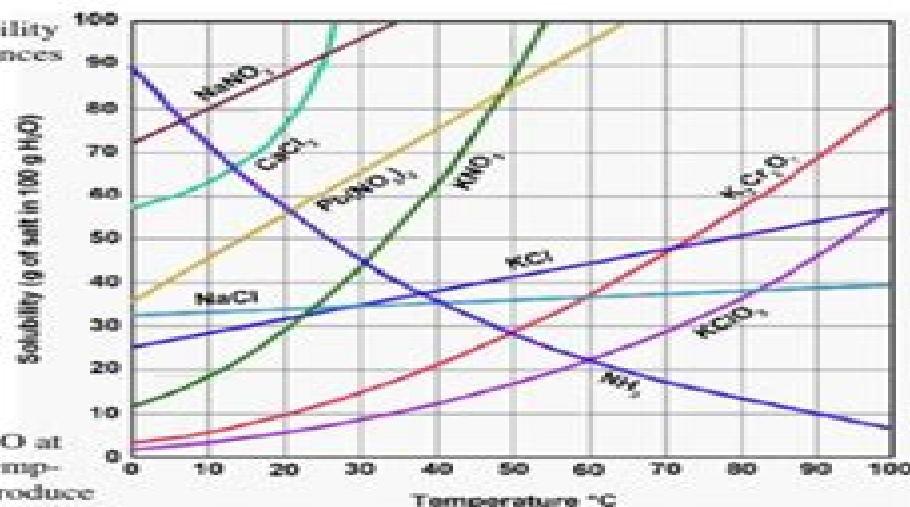
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Solubility Curves

The following graph shows solubility curves. Study the graph then answer the following questions.

1. What relationship exists between solubility and temperature for most of the substances shown?
- 2a. What is the exception?
b. What general principle accounts for this exception?
- 3a. Approximately how many grams of NaNO_3 will dissolve in 100 g water at 10 °C?
b. How many grams will dissolve at 34 °C?
4. How many grams of KCl will dissolve in 1 liter of H_2O at 50 °C?
5. 58.0 g of $\text{K}_2\text{Cr}_2\text{O}_7$ is added to 100 g H_2O at 0 °C. With constant stirring, to what temperature must the solution be raised to produce a saturated solution with no solid $\text{K}_2\text{Cr}_2\text{O}_7$ remaining?
6. A saturated solution of KClO_3 was made with 300 g of H_2O at 34 °C. How much KClO_3 could be recovered by evaporating the solution to dryness?
7. 500 g of water are used to make a saturated solution of KCl at 15 °C. How many more grams of KCl could be dissolved if the temperature were raised to 78 °C?
8. A saturated solution of KNO_3 in 200 g of H_2O at 51 °C is cooled to 28 °C. How much KNO_3 will precipitate out of solution?



1. _____

- 2b. _____

1. answer on left
- 2a. _____
b. answer on left
- 3a. _____
b. _____
4. _____
5. _____
6. _____
7. _____
8. _____

Solubility Of A Salt Lab Answer Key

**David Sang, Peter Ellis, Derek
McMonagle**



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