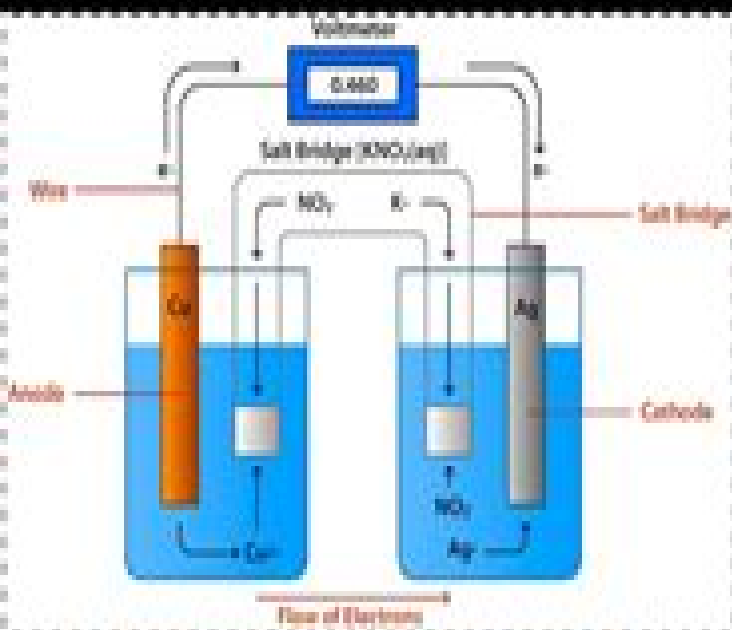
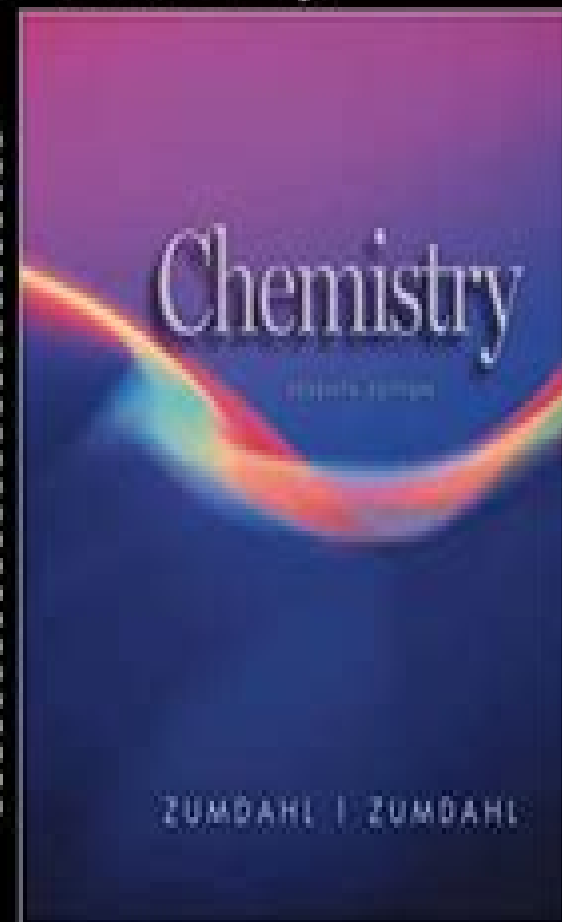


# ZUMDAHL CHEMISTRY 7TH ED.

## CHAPTER 17 (ELECTROCHEM)



| Oxidizing Agent                                                                                          | Reducing Agent                                                                                           | Standard Potential (V) |
|----------------------------------------------------------------------------------------------------------|----------------------------------------------------------------------------------------------------------|------------------------|
| F <sub>2</sub> + 2e <sup>-</sup> → 2F <sup>-</sup>                                                       | 2F <sup>-</sup> → F <sub>2</sub> + 2e <sup>-</sup>                                                       | 2.87                   |
| H <sub>2</sub> O <sub>2</sub> + 2H <sup>+</sup> + 2e <sup>-</sup> → 2H <sub>2</sub> O                    | 2H <sub>2</sub> O → H <sub>2</sub> O <sub>2</sub> + 2H <sup>+</sup> + 2e <sup>-</sup>                    | 1.78                   |
| MnO <sub>4</sub> <sup>-</sup> + 8H <sup>+</sup> + 5e <sup>-</sup> → Mn <sup>2+</sup> + 4H <sub>2</sub> O | Mn <sup>2+</sup> + 4H <sub>2</sub> O → MnO <sub>4</sub> <sup>-</sup> + 8H <sup>+</sup> + 5e <sup>-</sup> | 1.51                   |
| Ag <sup>+</sup> + e <sup>-</sup> → Ag                                                                    | Ag → Ag <sup>+</sup> + e <sup>-</sup>                                                                    | 1.50                   |
| Cl <sub>2</sub> + 2e <sup>-</sup> → 2Cl <sup>-</sup>                                                     | 2Cl <sup>-</sup> → Cl <sub>2</sub> + 2e <sup>-</sup>                                                     | 1.36                   |
| O <sub>2</sub> + 4H <sup>+</sup> + 4e <sup>-</sup> → 2H <sub>2</sub> O                                   | 2H <sub>2</sub> O → O <sub>2</sub> + 4H <sup>+</sup> + 4e <sup>-</sup>                                   | 1.23                   |
| O <sub>3</sub> /O <sub>2</sub> + 2H <sup>+</sup> + 2e <sup>-</sup> → O <sub>2</sub> + H <sub>2</sub> O   | O <sub>2</sub> + H <sub>2</sub> O → O <sub>3</sub> /O <sub>2</sub> + 2H <sup>+</sup> + 2e <sup>-</sup>   | 1.23                   |
| Br <sub>2</sub> + 2e <sup>-</sup> → 2Br <sup>-</sup>                                                     | 2Br <sup>-</sup> → Br <sub>2</sub> + 2e <sup>-</sup>                                                     | 1.07                   |
| NO <sub>3</sub> <sup>-</sup> + 4H <sup>+</sup> + 3e <sup>-</sup> → NO + 2H <sub>2</sub> O                | NO + 2H <sub>2</sub> O → NO <sub>3</sub> <sup>-</sup> + 4H <sup>+</sup> + 3e <sup>-</sup>                | 0.96                   |
| Ag <sup>+</sup> + e <sup>-</sup> → Ag                                                                    | Ag → Ag <sup>+</sup> + e <sup>-</sup>                                                                    | 0.80                   |
| I <sub>2</sub> + 2e <sup>-</sup> → 2I <sup>-</sup>                                                       | 2I <sup>-</sup> → I <sub>2</sub> + 2e <sup>-</sup>                                                       | 0.54                   |
| Cu <sup>2+</sup> + 2e <sup>-</sup> → Cu                                                                  | Cu → Cu <sup>2+</sup> + 2e <sup>-</sup>                                                                  | 0.52                   |
| O <sub>2</sub> + 2H <sub>2</sub> O + 4e <sup>-</sup> → 4OH <sup>-</sup>                                  | 4OH <sup>-</sup> → O <sub>2</sub> + 2H <sub>2</sub> O + 4e <sup>-</sup>                                  | 0.40                   |
| Cu <sup>+</sup> + e <sup>-</sup> → Cu                                                                    | Cu → Cu <sup>+</sup> + e <sup>-</sup>                                                                    | 0.34                   |
| 2H <sub>2</sub> O <sub>2</sub> + 2e <sup>-</sup> → H <sub>2</sub> + 2H <sub>2</sub> O                    | H <sub>2</sub> + 2H <sub>2</sub> O → 2H <sub>2</sub> O <sub>2</sub> + 2e <sup>-</sup>                    | 0.00                   |
| Hg <sup>2+</sup> + 2e <sup>-</sup> → Hg                                                                  | Hg → Hg <sup>2+</sup> + 2e <sup>-</sup>                                                                  | -0.85                  |
| Sn <sup>4+</sup> + 2e <sup>-</sup> → Sn <sup>2+</sup>                                                    | Sn <sup>2+</sup> → Sn <sup>4+</sup> + 2e <sup>-</sup>                                                    | -0.14                  |
| Fe <sup>3+</sup> + e <sup>-</sup> → Fe <sup>2+</sup>                                                     | Fe <sup>2+</sup> → Fe <sup>3+</sup> + e <sup>-</sup>                                                     | -0.28                  |
| Fe <sup>3+</sup> + e <sup>-</sup> → Fe <sup>2+</sup>                                                     | Fe <sup>2+</sup> → Fe <sup>3+</sup> + e <sup>-</sup>                                                     | -0.40                  |
| Cr <sup>3+</sup> + e <sup>-</sup> → Cr <sup>2+</sup>                                                     | Cr <sup>2+</sup> → Cr <sup>3+</sup> + e <sup>-</sup>                                                     | -0.74                  |
| Zn <sup>2+</sup> + 2e <sup>-</sup> → Zn                                                                  | Zn → Zn <sup>2+</sup> + 2e <sup>-</sup>                                                                  | -0.76                  |
| 2H <sub>2</sub> O + 2e <sup>-</sup> → H <sub>2</sub> + 2OH <sup>-</sup>                                  | H <sub>2</sub> + 2OH <sup>-</sup> → 2H <sub>2</sub> O + 2e <sup>-</sup>                                  | -0.83                  |
| Mn <sup>2+</sup> + 2e <sup>-</sup> → Mn                                                                  | Mn → Mn <sup>2+</sup> + 2e <sup>-</sup>                                                                  | -1.18                  |
| Al <sup>3+</sup> + 3e <sup>-</sup> → Al                                                                  | Al → Al <sup>3+</sup> + 3e <sup>-</sup>                                                                  | -1.68                  |



# Zumdahl Chemistry Chapter 17

**Michal Rosen-Zvi**



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