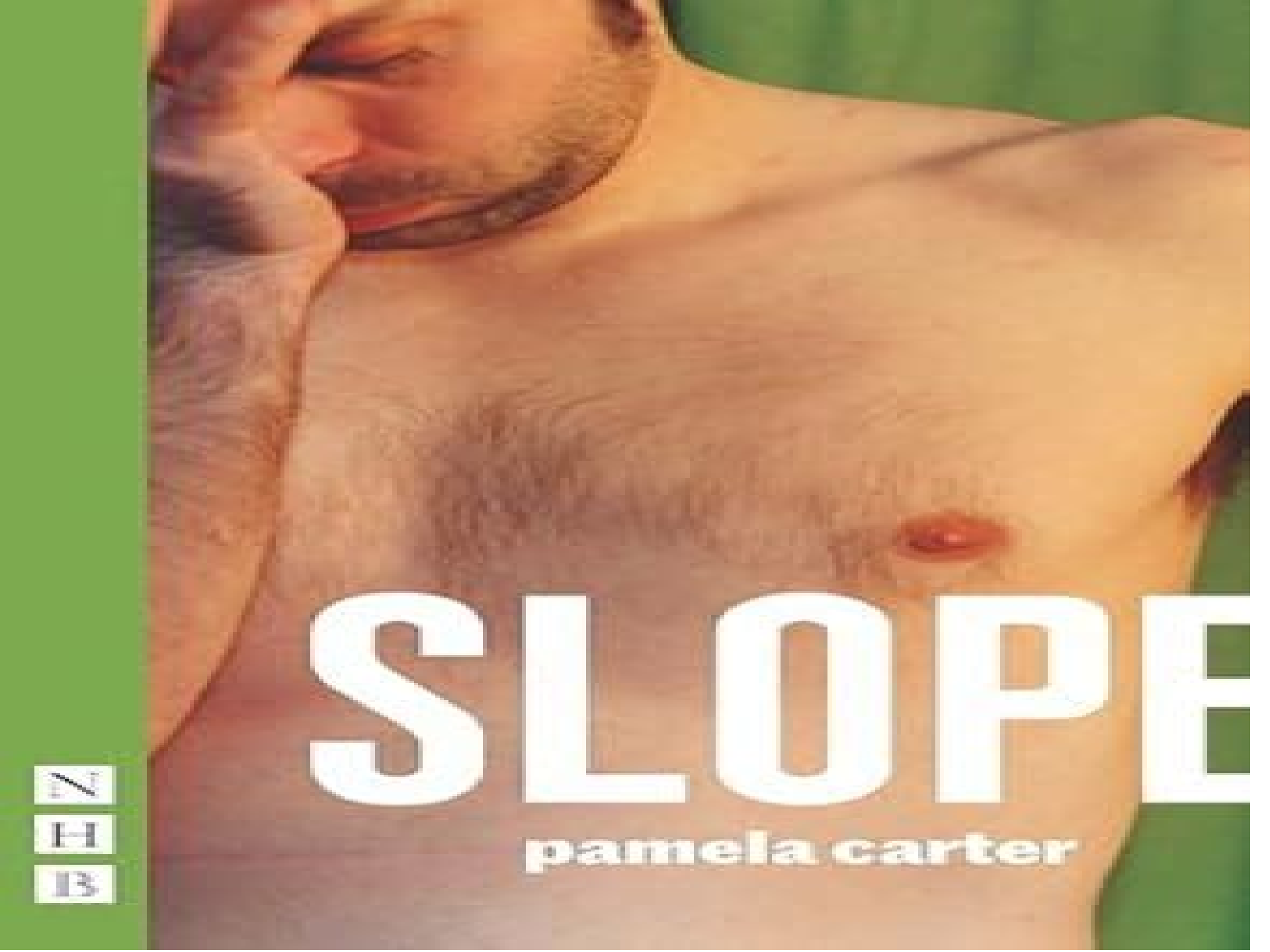




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 $\text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
 $2 \text{NaF} + \text{Cl}_2 \rightarrow 2 \text{H}_2 + 1 \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
 $\text{Pb} + \text{OH}^- \rightarrow 2 \text{HCl} \rightarrow 2 \text{H}_2\text{O} + 1 \text{PbCl}_2$
 $\text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + 1 \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 + 2 \text{O}_2 \rightarrow 1 \text{CO}_2 + 2 \text{H}_2\text{O} + \text{C}_3\text{H}_8$
 $5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O} + \text{C}_8\text{H}_{18}$
 $25 \text{O}_2 \rightarrow 16 \text{CO}_2 + 18 \text{H}_2\text{O}$

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 solid calcium carbonate is heated and decomposes to solid calcium oxide and carbon dioxide gas
 gaseous butane C_4H_{10} reacts with diatomic oxygen gas to yield gaseous carbon dioxide and water vapor

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 $1 \text{Na} + 3 \text{PO}_4 + 3 \text{KOH} + 3 \text{NaOH} + 1 \text{K} + 3 \text{PO}_4 + 1 \text{MgF}_2 + 1 \text{Li} + 2 \text{CO}_3 + 1 \text{MgCO}_3 + 2 \text{LiF} + 1 \text{P}_4 + 3 \text{O}_2 + 2 \text{P}_2\text{O}_3 + 2 \text{RbNO}_3 + 1 \text{BeF}_2 + 1 \text{BeNO}_3 + 2 \text{RbF}$

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 $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
 $2 \text{NaF} + \text{Cl}_2 \rightarrow 2 \text{H}_2 + 1 \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$
 $5 \text{Pb} + \text{OH}^- \rightarrow 2 \text{HCl} + 2 \text{H}_2\text{O} + 1 \text{PbCl}_2$
 $2 \text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + 1 \text{Al}_2\text{SO}_4 + 3 \text{CH}_4 + 2 \text{O}_2 \rightarrow 1 \text{CO}_2 + 2 \text{H}_2\text{O} + 8 \text{C}_3\text{H}_8 + 5 \text{O}_2 + 3 \text{CO}_2 + 4$

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 $1 \text{O}_2 + \text{NH}_3 \rightarrow \text{HNO}_3 + \text{H}_2\text{O}$
 $2 \text{O}_2 + 3 \text{H}_2\text{O} + 3 \text{H}$

2 o 2 h 2 o o 2 4 rb h 2 o rboh h 2 5 cl 2 kbr kcl br 2 6 n 2 h 2 nh 3 7 ndf 3 ca nd caf 2 8 ch 4 h 2 o co h 2 9 c 2 h 5 oh

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