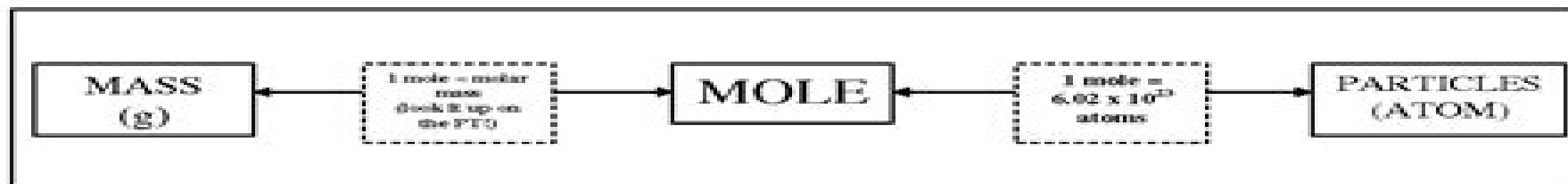


1 mole = 6.02×10^{23} atoms 1 mole = atomic mass (g)



Two Step Problems:

1. What is the mass of 2.25×10^{25} atoms of lead?

$$\frac{2.25 \times 10^{25} \text{ atoms of lead}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{207.2 \text{ g}}{1 \text{ mole}} = 7744.19 \text{ g} = 7740 \text{ g}$$

2. How many atoms are in 10.0g of gold?

$$\frac{10 \text{ g gold}}{197.0 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.06 \times 10^{22} \text{ atoms}$$

PRACTICE PROBLEMS:

1. How many moles are equal to 625g of copper?

$$\frac{625 \text{ g of copper}}{63.5 \text{ g Cu}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 9.77 \text{ mol Cu}$$

2. How many moles of barium are in a sample containing 4.25×10^{23} atoms of barium?

$$\frac{4.25 \times 10^{23} \text{ atoms of barium}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.706 \text{ mol}$$

3. Convert 2.35 moles of carbon to atoms.

$$\frac{2.35 \text{ moles}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 1.42 \times 10^{24} \text{ atoms}$$

4. How many atoms are in 4.0g of potassium?

$$\frac{4.0 \text{ g}}{39.1 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 6.2 \times 10^{22} \text{ atoms}$$

The Mole Chapter Worksheet

JR Anderson



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Table of Contents The Mole Chapter Worksheet

1. Understanding the eBook The Mole Chapter Worksheet
 - The Rise of Digital Reading The Mole Chapter Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying The Mole Chapter Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an The Mole Chapter Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from The Mole Chapter Worksheet
 - Personalized Recommendations
 - The Mole Chapter Worksheet User Reviews and Ratings
 - The Mole Chapter Worksheet and Bestseller Lists
5. Accessing The Mole Chapter Worksheet Free and Paid eBooks
 - The Mole Chapter Worksheet Public Domain eBooks
 - The Mole Chapter Worksheet eBook Subscription Services
 - The Mole Chapter Worksheet Budget-Friendly Options

6. Navigating The Mole Chapter Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - The Mole Chapter Worksheet Compatibility with Devices
 - The Mole Chapter Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of The Mole Chapter Worksheet
 - Highlighting and Note-Taking The Mole Chapter Worksheet
 - Interactive Elements The Mole Chapter Worksheet
8. Staying Engaged with The Mole Chapter Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers The Mole Chapter Worksheet
9. Balancing eBooks and Physical Books The Mole Chapter Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection The Mole Chapter Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine The Mole Chapter Worksheet
 - Setting Reading Goals The Mole Chapter Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of The Mole Chapter Worksheet
 - Fact-Checking eBook Content of The Mole Chapter Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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