Name.

Date:

Class

8.2

THE NATURE OF COVALENT BONDING

Section Review

Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used.
- Predict when two atoms are likely to be joined by a double or a triple covalent bond.
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

Vocabulary

- single covalent bond.
- structural formulas
- unestrained pairs

- double covalent bonds
- triple covalent bonds
- coordinate covalent bond.
- polyatomic ion.
- bond dissociation energy
- resonance structures

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase or number.

When atoms share electrons to gain the __1__ configuration 1.

of a noble gas, the bonds formed are __2__, A __3__ pair of 2.

valence electrons constitutes a __4__ covalent bond. Pairs of 3.

valence electrons that are not shared between atoms are called 4.

__5__, Sometimes two or three pairs of electrons may be shared 5.

to give __6__ covalent bonds. In some cases, only one of the 6.

atoms in a bond provides the pair of bonding electrons; this is a 7.

__7___ 8__ is required to break covalent bonds between 8.

atoms. The total energy required to break the bond between two 9.

covalently bonded atoms is known as the __9__. 10.

When it is possible to write two or more valid electron dot

formulas for a molecule or ion, each formula is referred to as a 10 ...

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Michael Rajnik

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