Date \_\_\_\_\_ Class

# THE NATURE OF COVALENT BONDING

# Section Review

# Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

#### Vocabulary

- · single covalent bond
- structural formulas
- unshared pairs

- double covalent bonds
- triple covalent bonds.
- coordinate covalent bond
- polyatomic ion
- bond dissociation energy
- resonance structures

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the \_\_\_\_1\_\_ configuration of a noble gas, the bonds formed are 2 A 3 pair of valence electrons constitutes a \_\_\_4 \_\_ covalent bond. Pairs of valence electrons that are not shared between atoms are called 5 Sometimes two or three pairs of electrons may be shared to give \_\_\_6\_\_ covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the \_\_\_9\_\_.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a \_\_

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2.	covalent	

- a shared

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# **The Nature Of Covalent Bonding 8 2 Answer Key**

**Rosina Ehmann** 

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