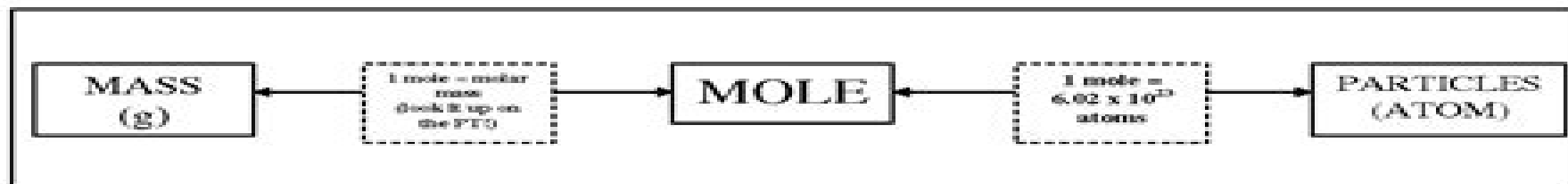


1 mole = 6.02×10^{23} atoms 1 mole = atomic mass (g)



Two Step Problems:

1. What is the mass of 2.25×10^{25} atoms of lead?

$$\frac{2.25 \times 10^{25} \text{ atoms of lead}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{207.2 \text{ g}}{1 \text{ mole}} = 7744.19 \text{ g} = 7740 \text{ g}$$

2. How many atoms are in 10.0g of gold?

$$\frac{10 \text{ g gold}}{197.0 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 3.06 \times 10^{22} \text{ atoms}$$

PRACTICE PROBLEMS:

1. How many moles are equal to 625g of copper?

$$\frac{625 \text{ g of copper}}{63.5 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 9.77 \text{ mol Cu}$$

2. How many moles of barium are in a sample containing 4.25×10^{23} atoms of barium?

$$\frac{4.25 \times 10^{23} \text{ atoms of barium}}{6.02 \times 10^{23} \text{ atoms}} \times \frac{1 \text{ mole}}{1 \text{ mole}} = 0.706 \text{ mol}$$

3. Convert 2.35 moles of carbon to atoms.

$$\frac{2.35 \text{ moles}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 1.42 \times 10^{24} \text{ atoms}$$

4. How many atoms are in 4.0g of potassium?

$$\frac{4.0 \text{ g}}{39.1 \text{ g}} \times \frac{1 \text{ mole}}{1 \text{ mole}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 6.2 \times 10^{22} \text{ atoms}$$

The Mole Chapter Answer Key

**Linda A. LaCharity, Candice K.
Kumagai, Barbara Bartz**



The Mole Chapter Answer Key:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises [The Practice of Chemistry](#) Donald J. Wink, Sharon Fetzner-Gislason, Sheila McNicholas, 2003-03

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