

Directions:

- You must bring a pencil and calculator to the exam.
- You will be given a periodic table.
- If you complete the entire review sheet by the day of your exam you will be eligible for the curve should one be applied.
- The answers to this review must be on a separate sheet of lined paper. You will first see the objective of each question followed by the question itself. The question is what you are providing the answer to.

Objectives and Questions:

Chapter 1: Thinking Like a Scientist

- 1) The labs you do help you process and analyze information. They also teach you how to design labs. Be aware of how lab design works. Think critically about the errors that can occur.
 - a. Why are we do need multiple trials? Constants? Controls? Variables? Hypotheses?
 - b. Define accuracy and precision. It is possible to collect data that is both accurate and precise? Is it possible to collect data that is precise but not accurate?
 - c. What can you do when you run a lab to eliminate errors? Please make sure to refer to the equipment, the variables, and the materials used for the lab.
 - d. Can theories be modified? How has the atomic theory been modified over time?

Chapter 2: Atoms

- 1) Know how to identify a compound as an ionic or covalent compound. Covalent compounds are also referred to as molecular compounds.
 - a. What types of elements are in binary ionic compounds? Covalent Compounds (also called binary molecular compounds)?
 - b. Circle the following formulas that represent ionic compounds. Put a box around those that are covalent/molecular.

FeO_2 NaCl BaCl_2 O_2 FeSO_4 H_2O_2 Fe_2O_3 $(\text{NH}_4)_2\text{CO}_3$ PH_3
- 2) Know how to write the names and formulas for ionic compounds, covalent compounds, acids and bases.
 - a. Write the names for the following compounds:
 - i. SnO_2
 - ii. H_2S
 - v. KOH
 - vi. $\text{Mg}(\text{CN})_2$
 - ii. $\text{Fe}_2(\text{PO}_4)_3$
 - iv. HNO_3
 - vi. CO
 - viii. HNO_2
 - b. Write the formulas for the following compounds:
 - i. Silicon dioxide
 - ii. Phosphoric acid
 - iii. Lithium sulfite
 - iv. Tricarbon octahydride
 - v. potassium fluoride
 - vi. sodium hypochlorite
 - vii. Hydrocyanic acid
 - viii. Iron (II) oxide
 - c. Why do certain compounds require a roman numeral? What block of elements requires them?
- 3) Be able to predict the products of an acid and base reaction. Know what pH is and how to classify a solution as acidic, basic or neutral. Be familiar with common methods of pH detection.
 - a. Based on the following pH's, classify the solutions as acidic, basic, or neutral:
 - i. $\text{pH} = 2$
 - ii. $\text{pH} = 5.8$
 - iii. $\text{pH} = 7$
 - b. List two properties each for acids and bases. (Examples include taste, what ion they create, feel, and ability to react with a metal.)
 - c. What is more acidic, a solution with a pH of 2 or 5? Which is more basic, a pH of 8 or 14?

Utica Community Schools Chemistry Final Review

Yicheng Fang



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Resources in Education ,1993-10 *Library Journal* Karl Brown,Melvil Dewey,Frederick Leypoldt,Bertine Emma Weston,Helen E. Wessels,1909 Includes beginning Sept 15 1954 and on the 15th of each month Sept May a special section School library journal ISSN 0000 0035 called Juniorlibraries 1954 May 1961 Issued also separately *Research Awards Index* ,1977 **Yale Alumni Weekly** ,1907 The Yale Alumni Weekly ,1908 Alumni News ,1929 The Nation ,1897
Educational Film Magazine ,1921 Power ,1929 *The College Courant* ,1874 **Who's who in America** ,1926

Onshore Unconventional Hydrocarbon Development National Academies of Sciences, Engineering, and Medicine,Division on Earth and Life Studies,Water Science and Technology Board,Board on Earth Sciences and Resources,Roundtable on Unconventional Hydrocarbon Development,2018-08-23 Oil and gas well completion and stimulation technologies to develop unconventional hydrocarbon resources in the United States have evolved over the past several decades particularly in relation to the development of shale oil and shale gas Shale oil and shale gas resources and the technology associated with their production are often termed unconventional because the oil and gas trapped inside the shale or other low permeability rock formation cannot be extracted using conventional technologies Since about 2005 the application of these technologies to fields in the U S have helped produce natural gas and oil in volumes that allowed the country to reduce its crude oil imports by more than 50% and to become a net natural gas exporter The regional and national economic and energy advances gained through production and use of these resources have been accompanied however by rapid expansion of the infrastructure associated with the development of these fields and public concern over the impacts to surface and groundwater air land and communities where the resources are extracted The intent of the first day of the workshop of the National Academies of Sciences Engineering and Medicine s Roundtable on Unconventional Hydrocarbon Development was to discuss onshore unconventional hydrocarbon development in the context of potential environmental impacts and the ways in which the risks of these kinds of impacts can be managed Specifically the workshop sought to examine the lifecycle development of these fields including decommissioning and reclamation of wells and related surface and pipeline infrastructure and the approaches from industry practice scientific research and regulation that could help to ensure management of the operations in ways that minimize impacts to the environment throughout their active lifetimes and after operations have ceased This publication summarizes the presentations and discussions from the workshop

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