#### Chapter 1

Given the atomic mass of uranium (<sup>238</sup>U) is 238.0289 u and that it has a nuclear radius of 7.5×10° <sup>15</sup> m, calculate the density of the nucleus of <sup>258</sup>U. Hence, estimate the mass of a fuel pellet of diameter 10 mm and height 8 mm if it were possible for it to be comprised entirely of uranium with the composition of its nucleus.

#### Answer

Assuming the <sup>138</sup>U nucleus is spherical, the volume V is:

$$V = \frac{4}{3}\pi r^2 = \frac{4}{3}\pi (7.5 \times 10^{-15})^2 \sim 1.8 \times 10^{-42} \text{ m}^2$$

The mass mu of 1 atom of 250U is:

$$m_U = 238.0289 \times 1.66 \times 10^{-27} = 4.0 \times 10^{-28} \text{ kg}$$

Hence, the density p is as follows:

$$\rho = \frac{4.0 \times 10^{-25}}{1.8 \times 10^{-42}} = 2.2 \times 10^{17} \text{ kg m}^{-3}$$

Thus the mass of the pellet would be:

$$m = 2.2 \times 10^{17} \times \pi (5 \times 10^{-3})^{2} \times 8 \times 10^{-3} = 1.4 \times 10^{11} \text{ kg}$$

Caesium-137 decays via β decay to an excited state of barium-137 which decays subsequently
via the emission of a single γ ray having an energy of 662 keV.
 Convert this energy to joules. How many <sup>137</sup>Cs γ rays would need to be absorbed by 1 kg of
water to raise its temperature by 1 °C if its specific heat capacity is 4.2 kJ kg<sup>1</sup> °C<sup>1</sup>?

#### Andrwer

1 keV = 
$$1.6 \times 10^{-16}$$
 J  
 $E_y = 662 \times 1.6 \times 10^{-16}$   
=  $1.06 \times 10^{-19}$  J  
 $E = mc\Delta T$   
=  $1 \times 4200 \times 1$   
=  $4200$  J  
 $N_y = \frac{4200}{1.00 \times 10^{-11}} = 4.0 \times 10^{16}$  y rays

# **Solution Manual For Nuclear Reactor**

**Tobias Bleicker** 

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