

1

Assignment 1. Questions from chapters 1 and 2 of McMurry and Fay
Question numbers are from the fourth edition.

Chapter 1. Chemistry: Matter and Measurement

- 1.1 (a) Cd (b) Sb (c) Am
- 1.2 (a) silver (b) rhodium (c) rhenium (d) cesium (e) argon (f) arsenic
- 1.3 (a) Ti, metal (b) Te, semimetal (c) Se, nonmetal
(d) Sc, metal (e) At, semimetal (f) Ar, nonmetal
- 1.4 The three *A*-coinage metals[®] are copper (Cu), silver (Ag), and gold (Au).
- 1.5 (a) The decimal point must be shifted ten places to the right so the exponent is ≤ 10 . The result is 3.72×10^{-10} m.
(b) The decimal point must be shifted eleven places to the left so the exponent is ≤ 11 . The result is 1.5×10^{11} m.
- 1.6 (a) microgram (b) decimeter (c) picosecond
(d) kiloampere (e) millimole
- 1.7 $^{\circ}\text{C} = \frac{5}{9} \times (^{\circ}\text{F} - 32) = \frac{5}{9} \times (98.6 - 32) = 37.0^{\circ}\text{C}$
 $\text{K} = ^{\circ}\text{C} + 273.15 = 37.0 + 273.15 = 310.2 \text{ K}$
- 1.8 (a) $\text{K} = ^{\circ}\text{C} + 273.15 = 578 + 273.15 = 851.15 \text{ K} = 851 \text{ K}$
(b) $^{\circ}\text{F} = \left(\frac{9}{5} \times ^{\circ}\text{C}\right) + 32 = \left(\frac{9}{5} \times 158\right) + 32 = 316.4^{\circ}\text{F} = 316^{\circ}\text{F}$
(c) $^{\circ}\text{C} = \text{K} - 273.15 = 375 - 273.15 = 101.85^{\circ}\text{C} = 102^{\circ}\text{C}$
 $^{\circ}\text{F} = \left(\frac{9}{5} \times ^{\circ}\text{C}\right) + 32 = \left(\frac{9}{5} \times 101.85\right) + 32 = 215.33^{\circ}\text{F} = 215^{\circ}\text{F}$
- 1.9 $d = \frac{m}{V} = \frac{27.43 \text{ g}}{12.40 \text{ cm}^3} = 2.212 \text{ g/cm}^3$
- 1.10 $\text{volume} = 9.37 \text{ g} \times \frac{1 \text{ mL}}{1.483 \text{ g}} = 6.32 \text{ mL}$
- 1.11 The actual mass of the bottle and the acetone = $38.0015 \text{ g} + 0.7791 \text{ g} = 38.7806 \text{ g}$. The measured values are 38.7798 g , 38.7795 g , and 38.7801 g . These values are both close to each other and close to the actual mass. Therefore the results are both precise and accurate.

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1. Understanding the eBook Solution Manual Chemistry 4th Ed McMurry
 - The Rise of Digital Reading Solution Manual Chemistry 4th Ed McMurry
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Chemistry 4th Ed McMurry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Chemistry 4th Ed McMurry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Chemistry 4th Ed McMurry
 - Personalized Recommendations
 - Solution Manual Chemistry 4th Ed McMurry User Reviews and Ratings
 - Solution Manual Chemistry 4th Ed McMurry and Bestseller Lists

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 - Interactive Elements Solution Manual Chemistry 4th Ed McMurry
8. Staying Engaged with Solution Manual Chemistry 4th Ed McMurry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Chemistry 4th Ed McMurry
9. Balancing eBooks and Physical Books Solution Manual Chemistry 4th Ed McMurry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Chemistry 4th Ed McMurry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Manual Chemistry 4th Ed McMurry
 - Setting Reading Goals Solution Manual Chemistry 4th Ed McMurry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solution Manual Chemistry 4th Ed McMurry
 - Fact-Checking eBook Content of Solution Manual Chemistry 4th Ed McMurry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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