

Chapter 1

two compounds. Thus, the ratio of red atoms in the two compounds to two blue atoms in the two compounds, is three to one.

- 1.19 The ratio of red between the two compounds compared to fixed number of blue atoms is not one-to-one so they are different compounds. You could also compare the ratio of blue to red in each compound to demonstrate that these are two different compounds.
- 1.20 These two molecules do not have a different ratio of red in the two compounds compared to a fixed number of blue atoms. The ratio of red to one blue atom is two red to one blue. Therefore, they do not demonstrate the law of multiple proportions.
- 1.21 This may stand for the name of an element or for the name of one atom of an element.
- 1.22 The smallest particle that is representative of a particular element is the atom of that element. A molecule is a representative unit that is made up of two or more atoms linked together.
- 1.23 H_2 , hydrogen N_2 , nitrogen, O_2 , oxygen F_2 , fluorine
 Cl_2 , chlorine Br_2 , bromine I_2 , iodine
- 1.24 (a) nitrogen, N
(b) bromine, Br
(c) chlorine, Cl
(d) oxygen, O
- 1.25 (a) carbon, C
(b) hydrogen, H
(c) sulfur, S
(d) iodine, I
- 1.26 DNA contains hydrogen, carbon, nitrogen, oxygen, and phosphorus atoms.
- 1.27 A chemical reaction is balanced when there is the same number of each kind of atom on both the reactant and product side of the equation; and the total charges on both the reactant and product sides of the equation are the same. These conditions must be met due to the law of conservation of matter.
- 1.28 Reactants are the substances to the left of the arrow in a reaction that are present before the reaction begins. Products are the substances to the right of the arrow in a reaction and they are formed during the reaction and are present when the reaction is over.
- 1.29 (a) Magnesium reacts with oxygen to give (yield) magnesium oxide.
(b) The reactants are Mg and O_2 .
(c) The product is MgO.

Review Problems

- 1.30 3 Fe, 2 Al, 3 Si, 12 O
- 1.31 3 Ca, 5 Mg, 8 Si, 24 O, 2 H
- 1.32 $\text{KNaC}_6\text{H}_4\text{O}_6$
- 1.33 $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$

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