

CHAPTER 2

ATOMIC STRUCTURE AND INTERATOMIC BONDING

PROBLEM SOLUTIONS

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Fundamental Concepts

Electrons in Atoms

2.1 Cite the difference between atomic mass and atomic weight.

Solution

Atomic mass is the mass of an individual atom, whereas atomic weight is the average (weighted) of the atomic masses of an atom's naturally occurring isotopes.

2.2 Silicon has three naturally occurring isotopes: 92.23% of ^{28}Si , with an atomic weight of 27.9769 amu, 4.67% of ^{29}Si , with an atomic weight of 28.9765 amu, and 3.10% of ^{30}Si , with an atomic weight of 29.9737 amu. On the basis of these data, confirm that the average atomic weight of Si is 28.0855 amu.

Solution

The average atomic weight of silicon (A_{Si}) is computed by adding fraction-of-occurrence/atomic weight products for the three isotopes—i.e., using Equation 2.2. (Remember, fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus

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