

Study Guide

Key Concepts

- Concentrations can be measured qualitatively and quantitatively.
- Molarity is the number of moles of solute dissolved per liter of solution.

$$\text{molarity (M)} = \frac{\text{moles of solute}}{\text{liters of solution}}$$

- Molality is the ratio of the number of moles of solute dissolved in 1 kg of solvent.

$$\text{molality (m)} = \frac{\text{moles of solute}}{\text{kg of solvent}}$$



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formulate approaches Equations that appear in various chapters are developed not through lengthy derivations but by more of an intuitive approach The equations are presented in their conceptual form rather than a separate convenient form applicable to each clinic situation This method is used to demonstrate how you can apply the initial conditions to the properties of the drug patient and or route of administration rather than memorizing each variation of the basic equation The author defines pertinent pharmacokinetic terms as well as kinetic processes and classical modeling relevant to dosing adjustments Examples are included within each chapter that emphasize an understanding of the concepts Pharmacokinetic Principles of Dosing Adjustments was written for practitioners who operate in a setting that requires careful consideration to dosing parameters and in particular with patients that require constant monitoring of therapeutic outcomes including dosing adjustments Based on the introductory course in pharmacokinetics taught by Dr Schoenwald for the past twenty years this book is intended as a review and resource for practicing pharmacists *The Gospel of Education* Martin

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