

Name : _____ Date : _____

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Solubility Rules Practice Worksheet

1. Classify each substance as being soluble or insoluble in water.

I. $\text{Mg}(\text{PO}_4)_2$ - _____

VI. KBr - _____

II. KOH - _____

VII. $\text{Pb}(\text{CO}_3)_2$ - _____

III. NiCl_2 - _____

VIII. PbI_2 - _____

IV. NH_4OH - _____

IX. BaSO_4 - _____

V. Hg_2SO_4 - _____

X. NiCl_2 - _____

2. Show the ions that formed the following compounds:

I. $\text{Zn}_3(\text{PO}_4)_2$

II. Al_2S_3

III. Iron (III) sulfide

IV. Ammonium cyanide

3. Form 4 water-soluble compounds by combining ions from the ions below:

Cl^-

CO_3^{2-}

PO_4^{3-}

Li^+

Sr^{2+}

4. Identify the precipitate in the following reaction. Circle the correct answer.



5. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$\text{Zn}_3(\text{PO}_4)_2$		
	Silver bromide	
KNO_3		
	Aluminum sulfide	
	Silver acetate	

Solubility Practice Problems

**John Olmsted, Greg Williams, Robert C.
Burk**



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