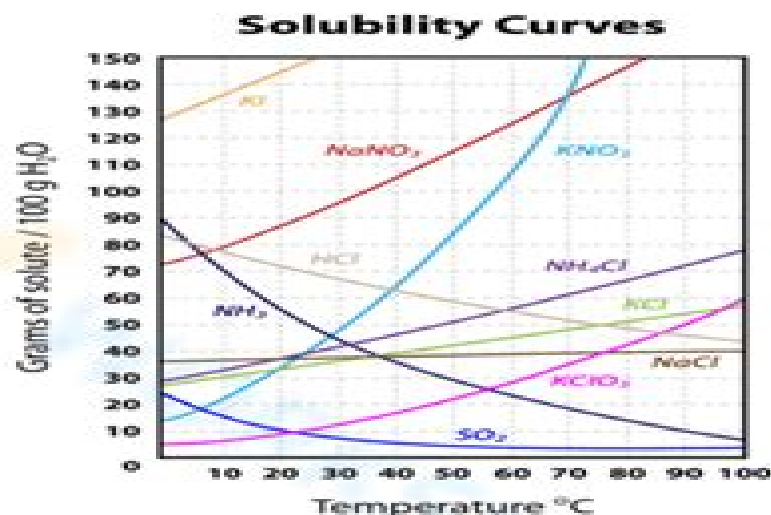


WS 6.1: Interpreting Solubility Curves

Directions: Use the Solubility Curve Table below to answer all questions. So you do not have to keep flipping your paper over you might want to put the solubility curve on your phone or computer screen when answering the questions on the back. All questions assume the amount of water is 100 mL.



- What is the solvent in this Solubility Curve?
water
- Saturation points are given per how much solvent?
100 ml or 100 g
- What is the saturation point of NaNO₃ at 10°C, 40°C, & 80°C?
@ 10°C = ~79 g
@ 40°C = ~105 g
@ 80°C = ~146 g
- How much NH₃ can you dissolve in water at 10°C, 30°C, & 90°C?
@ 10°C = ~70 g
@ 30°C = ~45 g
@ 90°C = ~10 g
- What is the maximum amount of KCl that can be dissolved in water at 20°C, 50°C, & 90°C?
@ 20°C = ~33 g
@ 50°C = ~42 g
@ 90°C = ~53 g
- If you want to make a saturated solution of NaCl how many grams would you dissolve in 100 grams of water @ 80°C?
About 40 grams
- Except for KI What substance is the most soluble in water at 20°C?
NaNO₃
- Except for KI, what substance is the most soluble in water at 72°C?
KNO₃
- Look up The boiling points for SO₂, HCl and NH₃ and write them below:
B.P. SO₂ = -10°C
B.P. HCl = -85°C
B.P. NH₃ = -33°C
- Based on the boiling points you looked up, are SO₂, HCl & NH₃ solids, liquids, or gases at 0°C? (and anywhere else on the solubility curve)
gases

Solubility Curves Answers Chemistry If8766

Atherton Seidell



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Organic Compounds Atherton Seidell,1928 **Developments and Applications in Solubility** Trevor M
Letcher,2007-10-31 Solubility is fundamental to most areas of chemistry and is one of the most basic of thermodynamic
properties It underlies most industrial processes Bringing together the latest developments and ideas Developments and
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