

## SNC1D3 – Solubility Curve worksheet SOLUTIONS

For questions 1-5 the amount of solute is given, and a temperature is stated. If all the solute could be dissolved in 100 grams of water at the stated temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 60g KCl at 70°C **supersaturated [above line]**
- 90g KNO<sub>3</sub> at 60°C **unsaturated [below line]**
- 110g NaNO<sub>3</sub> at 45°C **saturated [on line]**
- 10g KClO<sub>3</sub> at 10°C **supersaturated [above line]**
- 60g NH<sub>4</sub>Cl at 70°C **unsaturated [below line]**

For questions 6-10 a solute and a temperature are given. Tell how many grams of each solute must be added to 100 grams of water to form a **saturated solution** (on the line) at that given temperature

- NaNO<sub>3</sub> at 30°C **96g**
- KClO<sub>3</sub> at 70°C **36g**
- KNO<sub>3</sub> at 45°C **75g**
- KCl at 40°C **39g**
- NaCl at 90°C **40g**

For questions 11-13, tell which solution is more concentrated (more solute dissolved)

- At 50°C - **A) a saturated solution of KNO<sub>3</sub> [85g]**  
B) a saturated solution of NH<sub>4</sub>Cl [52g]
- At 50°C - **A) a saturated solution of KNO<sub>3</sub> [85g]**  
B) an unsaturated solution of NaNO<sub>3</sub> consisting of 100g of the solute dissolved in 100g of water.
- At 50°C - **A) a saturated solution of NaNO<sub>3</sub> [116g]**  
B) a supersaturated solution of NH<sub>4</sub>Cl consisting of 60g of the solute dissolved in 100g of water.

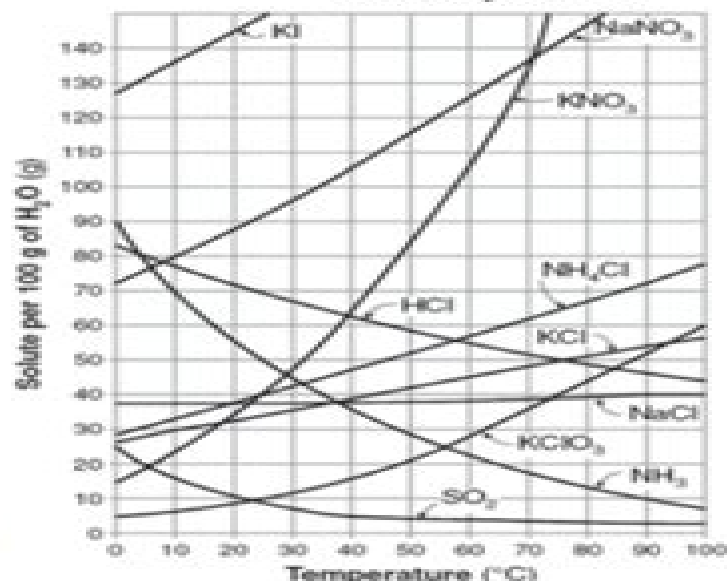
- Which compound is most soluble at 40°C? **KI**
- Which compound is least soluble at 90°C? **SO<sub>2</sub>**
- How many grams of KNO<sub>3</sub> will dissolve in 100g of water at 60°C? **108g**
- If 100g of a saturated solution of KNO<sub>3</sub> at 70°C is cooled to 20°C, how much KNO<sub>3</sub> will crystallize (in grams) **100g**

- If 200g of a saturated solution of KClO<sub>3</sub> at 50°C were evaporated to dryness, what would be left in the beaker?  
**42g KClO<sub>3</sub>**

OR.  $\frac{1135g}{135g} = \frac{x}{35g} \Rightarrow 100x = 21 \times 200 \Rightarrow 100x = 4200 \Rightarrow x = \frac{4200}{100} \Rightarrow x = 42$

- How much KCl will dissolve in 200g of water at 80°C **102g [51g x 2]**
- Which is more soluble at 30°C? KNO<sub>3</sub> or NaNO<sub>3</sub>? **NaNO<sub>3</sub>**
- How much NH<sub>4</sub>Cl will dissolve in 50g of water at 60°C? **29g**  
 $\frac{58g}{100g} = \frac{x}{50g} \Rightarrow 100x = 50 \times 58 = 100x = 2900 \Rightarrow x = \frac{2900}{100} \Rightarrow x = 29$  OR. **[58g/2 in 100g of water/2 = 29]**
- A saturated solution of NH<sub>4</sub>Cl was made at 50°C in 100g of water. If the temperature is raised to 90°C, how much more NH<sub>4</sub>Cl must be added to be saturated? **20g [52g in 50°C and 72g in 90°C  $\Rightarrow 72-52 = 20$ ]**
- A saturated solution of KClO<sub>3</sub> at 70°C in 200g of water is cooled to 18°C. How much solute will precipitate out? **56g**  
**[36g at 70°C in 100g of water so 36x2=72 at 70°C in 200g of water  $\Rightarrow$  8g at 18°C in 100g of water, so 8x2=16 at 18°C in 200g of water]**  
**[72 - 16 = 56]**
- How much more SO<sub>2</sub> will dissolve in 50g of water at 10°C than at 40°C? **5.5g**  
**[16/2 at 10°C = 8g  $\Rightarrow$  5g/2 at 40°C = 2.5g] [8 - 2.5 = 5.5]**

### Solubility Curves



# Solubility Curve Practice Problems 1 Key

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