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Chapter 6 – Quantities in Chemical Reactions

- 6.1 (a) stoichiometry; (b) heat; (c) endothermic reaction; (d) specific heat; (e) actual yield
- 6.2 (a) limiting reactant; (b) exothermic reaction; (c) theoretical yield; (d) law of conservation of energy; (e) percent yield
- 6.3 (a) The balanced chemical equation for the metabolism of glucose is:
$$\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$$
- (b) The coefficients of the balanced chemical equation give the proportions of the reactants and products in a reaction. From the balanced chemical equation, we see that the proportion of carbon dioxide to glucose is 6 molecules of CO_2 to 1 molecule of $\text{C}_6\text{H}_{12}\text{O}_6$. When 12 molecules of $\text{C}_6\text{H}_{12}\text{O}_6$ react, 72 (12×6) molecules of CO_2 form:
$$12 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ molecules } \text{CO}_2}{1 \text{ molecule } \text{C}_6\text{H}_{12}\text{O}_6} = 72 \text{ molecules } \text{CO}_2$$
- (c) From the balanced equation we see that glucose reacts with oxygen in a ratio of 1 $\text{C}_6\text{H}_{12}\text{O}_6$ molecule to 6 O_2 molecules. If 30 molecules of O_2 react, then 5 $\text{C}_6\text{H}_{12}\text{O}_6$ molecules will react:
$$30 \text{ molecules } \text{O}_2 \times \frac{1 \text{ molecule } \text{C}_6\text{H}_{12}\text{O}_6}{6 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6} = 5 \text{ molecules } \text{C}_6\text{H}_{12}\text{O}_6$$
- 6.4 (a) The balanced equation for the reaction is:
$$2\text{CH}_3\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 4\text{H}_2\text{O}$$
- (b) The coefficients of the balanced chemical equation give the proportions of the reactants and products in a reaction. From the balanced chemical equation, we see that 4 molecules of water react for every 2 molecules of methanol that react (a 2:1 ratio). If 10 molecules of methanol react, we expect 20 molecules of water to form:
$$10 \text{ molecules } \text{CH}_3\text{OH} \times \frac{4 \text{ molecules } \text{H}_2\text{O}}{2 \text{ molecules } \text{CH}_3\text{OH}} = 20 \text{ molecules } \text{H}_2\text{O}$$

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PT Brinkman



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