

Name Tanya Benedict

Date 5/19/24

Percent Yield Practice Problems

Directions: Determine the percent yield for each scenario in the problems below. You may need to determine the theoretical yield using stoichiometry before determining the percent yield.

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100$$

1. Determine the percent yield if a reaction yielded 23.5 grams of water vapor. The theoretical yield of the reaction was 32.0 grams of water vapor.

$$\frac{23.5 \text{ g}}{32 \text{ g}} \times 100 = 73.4\%$$

2. A student performs an experiment and generates 5.0 grams of hydrogen gas. They calculate the theoretical yield of their reaction and determine it should have generated 7.5 grams of hydrogen gas. What was the percent yield for the student's experiment?

$$\frac{5 \text{ g}}{7.5} \times 100 = 66.7\%$$

3. An industrial-grade reaction generates 4,500 kg of product. The company anticipated a yield of 5,500 kg. What is the percent yield for this reaction?

$$\frac{4500 \text{ kg}}{5500 \text{ kg}} \times 100 = 81.8\%$$

4. A pharmaceutical drug company has a prescription drug that has a percent yield of 76%. If the company uses enough materials to theoretically generate 8,000 kg of product, how much will they actually produce?

$$\frac{76}{100} \times 8000 \text{ kg} = 6080 \text{ kg}$$

Skill Practice 44 Practice Percent Yield

**Pam Jarvis, Jane George, Wendy
Holland, Jonathan Doherty**



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