

Worksheet 13 – Intermolecular Forces

Chemical bonds are **intramolecular** forces, which hold atoms together as molecules.

Intermolecular forces exist **between** atoms and molecules, determining their **physical** properties; solubility, boiling points, melting points.

Intermolecular forces (IMF) can be ranked qualitatively using Coulomb's Law:

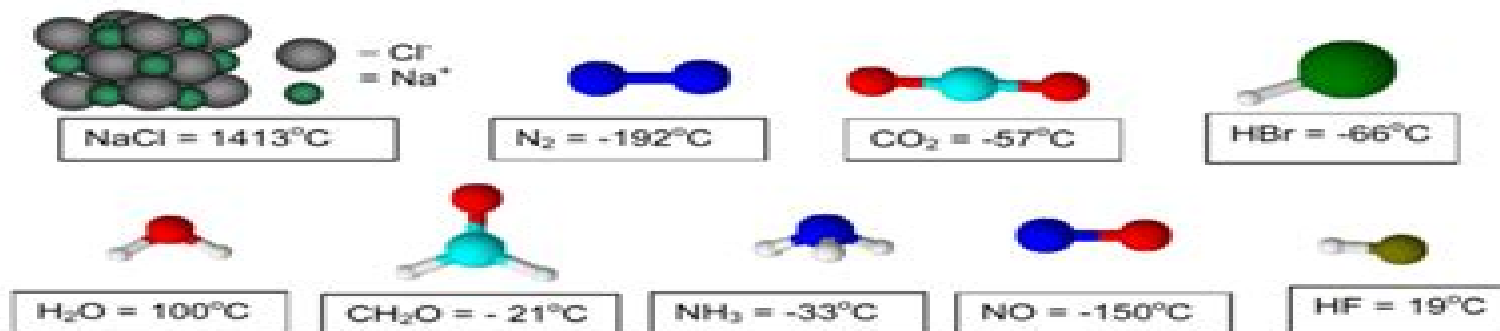
$$\text{force} \propto \frac{Q_1 Q_2}{r}$$

where Q_1 and Q_2 are charges and r is the distance between the charges.

The consequences are:

- the bigger the charges, the stronger the interaction
- the closer the charges, the stronger the interaction

Shown below are the structures and boiling points of nine chemical species. Use this information to answer the following questions.



1. Which species has the **highest** boiling point? _____

This is indicative of the **strongest IMF**.

- How does it differ from all of the other compounds?
- Look at Coulomb's law and identify each variable in the equation to the components of this species.
 Q_1 (positive charge) = _____ Q_2 (negative charge) = _____

This is called an _____ interaction (full positive and negative charges).

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Unveiling the Magic of Words: A Report on "**Skills Worksheet Concept Review Section Intermolecular Forces Answers**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Skills Worksheet Concept Review Section Intermolecular Forces Answers**," a mesmerizing literary masterpiece penned by a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound effect on the souls of its readers.

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