

## Moles to Particles Conversions

Name \_\_\_\_\_

Complete the following conversions. For FULL CREDIT, you must show ALL of your work!

1. Determine the number of atoms in 0.58 moles of Se.

$$0.58 \text{ mol Se} \times \frac{6.02 \times 10^{23} \text{ atoms Se}}{1 \text{ mol Se}} = 3.5 \times 10^{23} \text{ atoms Se}$$

2. How many moles are in  $6.80 \times 10^{24}$  formula units  $\text{Ba}(\text{NO}_3)_2$ ?

$$6.80 \times 10^{24} \text{ formula units Ba(NO}_3)_2 \times \frac{1 \text{ mol Ba(NO}_3)_2}{6.02 \times 10^{23} \text{ formula units Ba(NO}_3)_2} = 11.3 \text{ mol Ba(NO}_3)_2$$

3. Determine the number of molecules in 1.25 moles of  $\text{O}_2$  gas.

$$1.25 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} = 7.52 \times 10^{23} \text{ molecules O}_2$$

4. How many moles are in  $5.38 \times 10^{24}$  formula units  $\text{MgBr}_2$ ?

$$5.38 \times 10^{24} \text{ formula units MgBr}_2 \times \frac{1 \text{ mol MgBr}_2}{6.02 \times 10^{23} \text{ formula units MgBr}_2} = 8.94 \text{ mol MgBr}_2$$

5. Determine the number of formula units in 0.688 moles of  $\text{AgNO}_3$ .

$$0.688 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} = 4.14 \times 10^{23} \text{ formula units AgNO}_3$$

6. Convert  $8.46 \times 10^{24}$  molecules of ethane ( $\text{C}_2\text{H}_6$ ) to moles.

$$8.46 \times 10^{24} \text{ molecules C}_2\text{H}_6 \times \frac{1 \text{ mol C}_2\text{H}_6}{6.02 \times 10^{23} \text{ molecules C}_2\text{H}_6} = 14.1 \text{ mol C}_2\text{H}_6$$

7. Determine the number of formula units in 1.48 moles of  $\text{NaF}$ .

$$1.48 \text{ mol NaF} \times \frac{6.02 \times 10^{23} \text{ formula units NaF}}{1 \text{ mol NaF}} = 8.91 \times 10^{23} \text{ formula units NaF}$$

8. Convert  $8.2 \times 10^{23}$  molecules of water to moles.

$$8.2 \times 10^{23} \text{ molecules H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{6.02 \times 10^{23} \text{ molecules H}_2\text{O}} = 1.4 \text{ mol H}_2\text{O}$$

9. How many formula units are in 15 moles of  $\text{Li}_2\text{SO}_4$ ?

$$15 \text{ mol Li}_2\text{SO}_4 \times \frac{6.02 \times 10^{23} \text{ formula units Li}_2\text{SO}_4}{1 \text{ mol Li}_2\text{SO}_4} = 9.0 \times 10^{24} \text{ formula units Li}_2\text{SO}_4$$

10. Calculate the number of moles in  $7.2 \times 10^{23}$  atoms of sulfur.

$$7.2 \times 10^{23} \text{ atoms S} \times \frac{1 \text{ mol S}}{6.02 \times 10^{23} \text{ atoms S}} = 0.12 \text{ mol S}$$

\*Answer rounded using significant figures even/odd rounding rules.

# Skills Practice 31 Moles And Formulas

**J Spring**



## Skills Practice 31 Moles And Formulas:

**Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

*The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

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