

Name: _____

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Tribhuvan University
Institute of Science and Technology
B.Sc. Entrance Examination
Model question

Full marks: 100

Time: 2 hrs

Attempt all questions.

Group A (Chemistry)

(25 x 1 = 25)

- Which of the following elements does not contain neutron?
 - H
 - He
 - Be
 - Li
- The oxidation number of Nitrogen is - 1 in
 - NH₃
 - NH₂OH
 - N₂
 - N₂O
- The number of unpaired electrons in oxygen is
 - 2
 - 3
 - 4
 - 6
- The unit of radioactivity is
 - Fajan
 - Bohr
 - Curie
 - Debye
- Which of the following compounds gives positive AgNO₃ test?
 - CCl₄
 - CH₃Cl
 - CHCl₃
 - NaCl
- 'At same temperature and pressure equal volume of all gases has same number of molecules' this statement belongs to
 - Boyle's law
 - Avogadro's law
 - Charles' law
 - Graham's law
- A gas diffuses 4 times as fast as O₂. The molecular weight of the gas is
 - 4..2
 - 8
 - 16
 - 2
- In a solution containing 1 mole alcohol and 4 mole water, the mole fraction of alcohol is
 - 1/4
 - 3/4
 - 4/5
 - 1/5

Thermodynamic Examination Bank Question Papers

Jianjun Gao



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reaction a redox reaction explain your answer $2\text{NaCl aq} + \text{Pb} \rightarrow \text{Pb}^{2+} + 2\text{Cl}^{-}$ $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ PbCl_2

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$
 $5\text{Co} + \text{I}_2 \rightarrow 5\text{Co}^{2+} + \text{I}^-$ 2 $\text{O}_5 + 5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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state and oxidation number concepts to identify redox reactions and to keep track of electrons

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reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ Sr^0 to Sr^{2+} oxidized reducing agent O_2 to O^{2-} reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ Li^0 to Li^{+} oxidized red ag
 S^0 to S^{2-} reduced ox ag $3\text{Cs} + \text{Br}_2$