

## Thermochemistry Review Part 1

**Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for,

|                  |                        |                               |
|------------------|------------------------|-------------------------------|
| energy           | temperature            | standard state                |
| potential energy | heat                   | standard heat of formation    |
| kinetic energy   | specific heat capacity | exothermic                    |
| enthalpy         | endothermic            | Law of conservation of Energy |
| thermal          | Hess's Law             | isolated, closed, open system |

1. The specific heat capacity of diamond is  $0.5050 \text{ J/g}^\circ\text{C}$ . How much energy is required to heat  $25.0 \text{ g}$  of diamond from  $10.5^\circ\text{C}$  to  $15.6^\circ\text{C}$ ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of  $50 \text{ g}$  and an initial temperature of  $90^\circ\text{C}$  is placed into  $100 \text{ mL}$  of water at a temperature of  $25^\circ\text{C}$ . The temperature of water rises to  $31.3^\circ\text{C}$ . Determine the specific heat capacity of aluminum.

$$Q \text{ lost (Al)} = Q \text{ gained (Water)}$$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

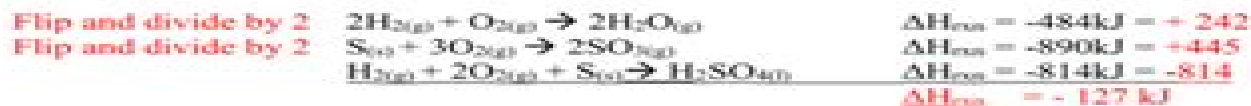
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by  $300 \text{ g}$  of methanol,  $\text{CH}_3\text{OH}$ , as it evaporates? Given: The molar heat of vaporization is  $35.3 \text{ kJ/mol}$ .

|                                  |                                                  |                                         |
|----------------------------------|--------------------------------------------------|-----------------------------------------|
| $q = ?$                          | $\Delta H = q/n$                                 | $n \text{ CH}_3\text{OH} = m/\text{MM}$ |
| $m = 300 \text{ g}$              | $q = H \times n$                                 | $n = 300 \text{ g} / 32 \text{ g/mol}$  |
| $\Delta H = 35.3 \text{ kJ/mol}$ | $q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$ | $n = 9.375 \text{ mol}$                 |
|                                  | $q = 332.81 \text{ kJ}$                          |                                         |

4. What is the molar enthalpy of the formation of  $1 \text{ mol H}_2\text{SO}_{4(l)}$  given the following information?



# Thermochemistry Test Answer Key

**Aditya Tripathi**



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**Nature Of Chemistry Volume - 1** Aditya Tripathi, *The Chemistry of the Metal-carbon Bond: The structure, preparation, thermochemistry, and characterization of organometallic compounds* F. R. Hartley, Saul Patai, 1982

*Pearson Edexcel A Level Chemistry (Year 1 and Year 2)* Andrew Hunt, Graham Curtis, Graham Hill, 2019-07-15 Develop and assess your students knowledge and skills throughout A level with worked examples practical assessment guidance and differentiated end of topic questions in this updated all in one textbook for Years 1 and 2 Combining everything your students need to know for the Pearson Edexcel A level Chemistry specification this revised textbook will Identify the level of your students understanding with diagnostic questions and a summary of prior knowledge at the start of the Student Book Provide support for all 16 required practicals with various activities and questions along with a Practical chapter covering procedural understanding and key ideas related to measurement Improve mathematical skills with plenty of worked examples including notes on methods to help explain the strategies for solving each type of problem Offer plenty of practice with Test yourself questions to help students assess their understanding and measure progress Encourage further reading and study with short passages of extension material Develop understanding with free online access to Test yourself answers and an extended glossary

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