

2. Complete the following using dimensional analysis:

a. 2.55 Joules to calories

to 0.60807 Joules

b. 150°C to J per Joule

to 150 Joules to Joules

c. 100 Joules to Joule

to 100 Joules

d. 100 Joule to Joule

to 100 Joules

3. Complete the following table for the following substances:

a. Diethylamine is a substance from 20.0°C to 100.0°C

b. Diethylamine is a substance from 100.0°C to 150.0°C

c. Diethylamine is a substance from 150.0°C to 200.0°C

Specific Heat Capacity: Problems

Complete the following table for the following substances. The table gives the specific heat capacity, the specific heat capacity, and the specific heat capacity.

a. The specific heat capacity of water is 4.18 J/kg°C. The specific heat capacity of water is 4.18 J/kg°C. The specific heat capacity of water is 4.18 J/kg°C.

b. A metal with specific heat capacity of 0.25 J/kg°C requires 10 J of heat to raise the temperature by 2°C. Calculate the mass of the metal? (20.0 g)

c. A substance requires 10 J of heat to raise the temperature by 1°C. The mass of the substance is 10 g. What is the specific heat capacity of the substance? (1.00 J/kg°C)

d. A metal with specific heat capacity of 0.25 J/kg°C and a mass of 10 g absorbs 10 J of heat. What will be the temperature change of the metal? (raised by 0.40°C)

Use the following information to answer questions 7-10:

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