

Thermochemistry Quiz 1

Questions 1-4 (1 pt. each); Questions 5-6 (2 pts. Each)

- _____ 1. An endothermic reaction causes the surroundings to
- warm up.
 - become acidic.
 - condense.
 - decrease in temperature.
 - release CO₂
- _____ 2. According to the first law of thermodynamics:
- energy is neither lost nor gained in any energy transformations.
 - perpetual motion is possible.
 - energy is conserved in quality but not in quantity.
 - energy is being created as time passes. We have more energy in this universe now than when time began.
- _____ 3. A calorimeter
- is equal to the molar enthalpy of a reaction.
 - is a dieting aid.
 - is an indicator of a spontaneous reaction.
 - is a device used to measure the transfer of heat energy.
 - is useful in measuring the amount of heat released by endothermic reactions.
- _____ 4. The specific heat of gold is 0.129 J/g•°C. What is the molar heat capacity of gold?
- 0.039 J/mol•°C
 - 0.129 J/mol•°C
 - 25.4 J/mol•°C
 - 39.0 kJ/mol•°C
 - 197 J/mol•°C
- _____ 5. Copper metal has a specific heat of 0.385 J/g•°C. Calculate the amount of heat required to raise the temperature of 22.8 g of Cu from 20.0°C to 875°C.
- 1.97×10^{-2} J
 - 1.0×10^{-2} J
 - 329 J
 - 7.51 kJ
 - 10.5 kJ
- _____ 6. The combustion of butane produces according to the equation
- $$2\text{C}_4\text{H}_{10}(\text{g}) + 13\text{O}_2(\text{g}) \rightarrow 8\text{CO}_2(\text{g}) + 10\text{H}_2\text{O}(\text{l}) \quad \Delta H^\circ_{\text{rxn}} = -5,314 \text{ kJ}$$
- How many grams of butane must be burned to release 6,375 kJ of heat?
- 1.20 g
 - 139 g
 - 0.0413 g
 - 69.7 g
 - 97.8 g

Questions 7-8: Label the following processes as either *exothermic* or *endothermic*. (1 pt. each)

7. Dry ice (solid CO₂) sits at room temperature and sublimates to CO₂ gas. _____
8. A block of solid aluminum is heated until it melts. _____

Thermochemistry Test Questions

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