

Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is $0.5050 \text{ J/g}^\circ\text{C}$. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of 50 g and an initial temperature of 90°C is placed into 100 mL of water at a temperature of 25°C . The temperature of water rises to 31.3°C . Determine the specific heat capacity of aluminum.

$$Q \text{ lost (Al)} = Q \text{ gained (Water)}$$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

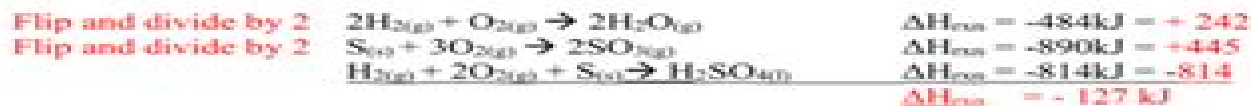
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by 300 g of methanol, CH_3OH , as it evaporates? Given: The molar heat of vaporization is 35.3 kJ/mol .

$q = ?$	$\Delta H = q/n$	$n \text{ CH}_3\text{OH} = m/\text{MM}$
$m = 300 \text{ g}$	$q = H \times n$	$n = 300 \text{ g} / 32 \text{ g/mol}$
$\Delta H = 35.3 \text{ kJ/mol}$	$q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$	$n = 9.375 \text{ mol}$
	$q = 332.81 \text{ kJ}$	

4. What is the molar enthalpy of the formation of $1 \text{ mol H}_2\text{SO}_{4(l)}$ given the following information?



Thermochemistry Review Sheet Answers

Patrick Vollmar



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The Electro-chemist and Metallurgist and Metallurgical Review ,1902 Robinson Chemistry Study Guide Robinson,1992 **Advances in Chemical Physics, Volume 19** Ilya Prigogine,Stuart A. Rice,2009-09-08 The Advances in Chemical Physics series provides the chemical physics and physical chemistry fields with a forum for critical authoritative evaluations of advances in every area of the discipline Filled with cutting edge research reported in a cohesive manner not found elsewhere in the literature each volume of the Advances in Chemical Physics series serves as the perfect supplement to any advanced graduate class devoted to the study of chemical physics

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