Thermochemistry Worksheet Key

- 1. How much heat, in calories, is released or absorbed if a 15.0 g sample of copper is converted from
 - (a) the solid to liquid state.

$$q = m \times \Delta H_{fus} = 15.0 g \times 206 \frac{J}{g} \times \frac{1 cal}{4.184 J} = 738.5 = 739 cal$$

So 739 cal is absorbed

(b) the gas to the liquid state.

$$q = m \times \Delta H_{con} = 15.0 g \times -4.7 \frac{kf}{g} \times \frac{1000 f}{1 kf} \times \frac{1 cal}{4.184 f} = -16.849$$

= -1.68 × 10⁴ cal or - 16.8 kcal

So 16.8 kcal is released.

For copper m.p. = 1084 °C; b.p. = 2927 °C; $\Delta H_{tim} = 206$ J/g, $\Delta H_{rap} = 4.7$ kJ/g

 How much heat is absorbed or released if a 12.5 g sample of aluminum changes temperature from 25.0 to 87.5 °C? The specific heat capacity of aluminum is 0.900 J/g.K.

$$T_{0mil} = 87.5 + 273.15 = 360.65 \text{ K}$$
 $T_{minil} = 25.0 + 273.15 = 298.15 \text{ K}$
 $q = m \times c \times \Delta T = 12.5 \text{ g} \times 0.900 \frac{J}{g.K} \times (360.65 - 298.15) K$
 $= 12.5 \text{ g} \times 0.900 \frac{J}{g.K} \times 62.5 \text{ K} = 703.1 = 703 J$

 What was the mass of a sample of water if the addition of 4.28 x 10³ joules raised the temperature of the sample from 22.0 °C to 34.5 °C?

T increases so endothermic so q > 0

$$q = m \times c \times \Delta T$$

$$so\ m = \frac{q}{c \times \Delta T} = \frac{4.28 \times 10^3 J}{4.184 \frac{J}{g \cdot K} \times (307.65 - 295.15) K} = \frac{4.28 \times 10^3 J}{4.184 \frac{J}{g \cdot K} \times 12.5 K} = 81.83 = 81.8 g$$

Thermochemistry Packet Key

Kaplan Test Prep

Thermochemistry Packet Key:

This book delves into Thermochemistry Packet Key. Thermochemistry Packet Key is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Thermochemistry Packet Key, encompassing both the fundamentals and more intricate discussions.

- 1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Thermochemistry Packet Key
 - Chapter 2: Essential Elements of Thermochemistry Packet Key
 - Chapter 3: Thermochemistry Packet Key in Everyday Life
 - Chapter 4: Thermochemistry Packet Key in Specific Contexts
 - ∘ Chapter 5: Conclusion
- 2. In chapter 1, the author will provide an overview of Thermochemistry Packet Key. This chapter will explore what Thermochemistry Packet Key is, why Thermochemistry Packet Key is vital, and how to effectively learn about Thermochemistry Packet Key.
- 3. In chapter 2, this book will delve into the foundational concepts of Thermochemistry Packet Key. The second chapter will elucidate the essential principles that need to be understood to grasp Thermochemistry Packet Key in its entirety.
- 4. In chapter 3, this book will examine the practical applications of Thermochemistry Packet Key in daily life. The third chapter will showcase real-world examples of how Thermochemistry Packet Key can be effectively utilized in everyday scenarios.
- 5. In chapter 4, this book will scrutinize the relevance of Thermochemistry Packet Key in specific contexts. The fourth chapter will explore how Thermochemistry Packet Key is applied in specialized fields, such as education, business, and technology.
- 6. In chapter 5, the author will draw a conclusion about Thermochemistry Packet Key. This chapter will summarize the key points that have been discussed throughout the book.
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