

Thermochemistry Review

Multiple Choice

Identify the choice that best completes the statement or answers the question.

Use the following information to answer the next _ questions.

Decaying matter may provide energy directly through combustion or indirectly through decomposition to form methane, which can then be burned.



- _____ 1. The burning of methane is an example of an
- exothermic reaction in which carbon is oxidized.
 - exothermic reaction in which carbon is reduced.
 - endothermic reaction in which carbon is oxidized.
 - endothermic reaction in which carbon is reduced.
- _____ 2. The molar enthalpy of combustion for methane is
- 560.5 kJ/mol
 - 710.1 kJ/mol
 - 802.3 kJ/mol
 - 951.9 kJ/mol
- _____ 3. An equation that represents an endothermic change is
- $\text{C}_2\text{H}_5\text{OH}(\text{l}) + 2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$
 - $2 \text{C}(\text{s}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g})$
 - $\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
 - $2 \text{C}(\text{s}) + 3 \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$
- _____ 4. Sulfuric acid is used to dissolve copper ore in many mining operations. Sulfuric acid may be produced on site by burning sulfur and dissolving the products in water. The overall equation for this reaction is
- $$\text{S}_8(\text{g}) + 12 \text{O}_2(\text{g}) + 8 \text{H}_2\text{O}(\text{l}) \rightarrow 8 \text{H}_2\text{SO}_4(\text{aq}) + 4225.6 \text{ kJ}$$
- In this process, what is the enthalpy of reaction for the production of one mole of sulfuric acid?
- +4225.6 kJ
 - +528.2 kJ
 - 528.2 kJ
 - 4225.6 kJ
- _____ 5. Which of the following statements is **false**?
- Light energy is converted into chemical energy during photosynthesis.
 - Chemical energy is converted into heat and light energy during the combustion of hydrocarbons.
 - Light energy is converted into chemical energy to produce a firefly's glow.
 - Chemical energy is converted into electrical energy as a voltaic cell discharges.

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Albert A Gayle



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