

Exercises

14.1-14.3 Temperature, Internal energy and states of matter

- Molecules (of the liquid) are in random / haphazard motion (AM)
 - Molecules (of liquid) are smaller than pollen grains.
 - increase the temperature (of the liquid)
- The sum of (the random distribution of) the KE and PE of (the) molecules
- No change in KE

Because temperature is constant (during melting)

PE of (the molecules) increases (during melting)

The internal energy increases.

14.4 Specific Heat Capacity

- 35 700 J
- 21 000 J/g°C
- 0.375 kg
- 66 J/s
- Electrical energy supplied = Heat taken in by water

$$I t V = m c \Delta \theta$$

$$2300 \times I \times 2 = 2.5 \times 4200 \times 80$$

$$I = 380 \text{ A [1]}$$
 - The kinetic energy of the particle is increasing. [1]
- $$E = 3 \times 60 = 180 \text{ J [1]}$$

Thermal Energy And Thermodynamics Test Review Answers

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Table of Contents Thermal Energy And Thermodynamics Test Review Answers

1. Understanding the eBook Thermal Energy And Thermodynamics Test Review Answers
 - The Rise of Digital Reading Thermal Energy And Thermodynamics Test Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermal Energy And Thermodynamics Test Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermal Energy And Thermodynamics Test Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermal Energy And Thermodynamics Test Review Answers
 - Personalized Recommendations
 - Thermal Energy And Thermodynamics Test Review Answers User Reviews and Ratings
 - Thermal Energy And Thermodynamics Test Review Answers and Bestseller Lists
5. Accessing Thermal Energy And Thermodynamics Test Review Answers Free and Paid eBooks
 - Thermal Energy And Thermodynamics Test Review Answers Public Domain eBooks
 - Thermal Energy And Thermodynamics Test Review Answers eBook Subscription Services
 - Thermal Energy And Thermodynamics Test Review Answers Budget-Friendly Options
6. Navigating Thermal Energy And Thermodynamics Test Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermal Energy And Thermodynamics Test Review Answers Compatibility with Devices
 - Thermal Energy And Thermodynamics Test Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermal Energy And Thermodynamics Test Review Answers
 - Highlighting and Note-Taking Thermal Energy And Thermodynamics Test Review Answers
 - Interactive Elements Thermal Energy And Thermodynamics Test Review Answers

8. Staying Engaged with Thermal Energy And Thermodynamics Test Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermal Energy And Thermodynamics Test Review Answers
9. Balancing eBooks and Physical Books Thermal Energy And Thermodynamics Test Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermal Energy And Thermodynamics Test Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermal Energy And Thermodynamics Test Review Answers
 - Setting Reading Goals Thermal Energy And Thermodynamics Test Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermal Energy And Thermodynamics Test Review Answers
 - Fact-Checking eBook Content of Thermal Energy And Thermodynamics Test Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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