Bonding & Molecular Structure

LEWIS STRUCTURES

Indicate the total # of VALENCE electrons for each molecule (VE's). Calculate the number of bonded electrons from the skeleton (BE's). Calculate the Remaining Electrons (RE's = VE's – BE's) and distribute the remaining electrons onto the skeletal structure. Finally, satisfy the octets by creating double and triple bonds. Your final structures should be given below and your scratch paper should clearly show your work.

F ₂ VE's = 14 BE's = 2 RE's = 12	H ₂ VE's = <u>2</u> BE's = <u>2</u> RE's = <u>0</u>	HF VE's = 8 BE's = 2 RE's = 6	$\frac{NH_3}{VE's} = \frac{3}{3}BE's = \frac{\zeta}{2}$ RE's = $\frac{\zeta}{2}$
٠ <u>Ë</u> - <u>F</u> :	н-н	H-Ë:	н-Й-н Н
Shape: linear	Shape: Linear C2H6	Shape: Linear	Shape: pyramiolal
VE's = 8 BE's = 8 RE's = 0	VE's = 14 BE's = 14 RE's = 6	VE's = 6 BE's = 6 RE's = 0	VE's = 6 BE's = 6 RE's = 0
н -¢-н Н Н	H - C - C - H	н - А\- н н	н-в-н
Shape: tetrahedral	Shape: tetrahectra Choth	Shape: trigonal	Shape: trigional
VE's = 12 BE's = 12 RE's = 0	VE's = 36 BE's = 12 RE's = 24	VE's = 10 BE's = 6 RE's = 4	VE's = 12 BE's = 4 RE's = 8
H C = C H	:Ë,c=c,Ë;	Ċ≡Ö	:Ö=Ö:
Shape around one C: trigonal planar	Shape around one C: trigonal plance	Shape: Lipear	Shapes

South Pasedena Bonding And Molecular Structure Worksheet Answers

Marcel A. Müller

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Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1980 A Pictorial Approach to Molecular Bondina John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as sociated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics. In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconcep tion that the ground rules governing bonding in molecules such as NH3 are somehow different from those which apply to aromatic systems such as C H 6 6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH for example Bonding and Molecular Structure Russell S. Drago, 1971 Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973 Structure of Molecules and the Chemical Bond Y. K. Syrkin, M. E. Dyatkina, 1950

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