

SOLUTION

From the point of view of thermodynamics, the maximum efficiency a Carnot engine can be depends on the lowest and highest temperatures it operates. The highest temperature in one cycle occurs at those points where the product $P \cdot V$ is the greatest. In this case, the greatest product takes place during the isothermal stage $P_1 - P_2$ (60 atm-liter). In a similar way, the lowest temperature occurs where the product $P \cdot V$ is the smallest, and in this problem is the isothermal stage $P_3 - P_4$ (8 atm-liter).

The highest temperature T_h is:

$$P \cdot V = n \cdot R \cdot T$$

$$60 \text{ atm-liter} = (0.3 \text{ mole})(0.0821 \text{ atm-liter/mole} \cdot ^\circ\text{K})(T_h)$$

Solving for the temperature T , the answer is:

$$T_h = 2,436 \text{ } ^\circ\text{K}$$

The lowest temperature T_l is:

$$P \cdot V = n \cdot R \cdot T$$

$$8 \text{ atm-liter} = (0.3 \text{ mole})(0.0821 \text{ atm-liter/mole} \cdot ^\circ\text{K})(T_l)$$

Solving for the temperature T , the answer is:

$$T_l = 324.6 \text{ } ^\circ\text{K}$$

$$\text{Maximum Efficiency} = (T_h - T_l)/T_h = 100$$

$$= (2,436 \text{ } ^\circ\text{K} - 324.6 \text{ } ^\circ\text{K})/2,436 \text{ } ^\circ\text{K} = 100$$

$$= 66.7\%$$

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Thermodynamics Problem And Solution

Camilla Rothe



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