
Practice Problem Set #1- Chapter 2 (Lec 2)**Problem 2.1**

A ship having a mass of 12000 kg is traveling at 60 km/h. How much work must be done to bring it to a stop?

Solution:

a) Using the first law of Thermo I:

$$\Delta U + \Delta KE + \Delta PE = Q + W$$

Assuming:

Isothermal: $\Delta U = 0$

Adiabatic: $Q = 0$

No elevation change: $\Delta PE = 0$

$$W = \Delta KE = \frac{1}{2} m(u_2^2 - u_1^2) = \frac{1}{2} \times 12000 \text{ kg} \left(0^2 - \left[60 \frac{\text{km}}{\text{h}} \times \frac{1000 \text{ m}}{1 \text{ km}} \times \frac{1 \text{ h}}{3600} \right]^2 \right)$$
$$W = -1.7 \times 10^6 \text{ J} = -1700 \text{ kJ}$$

Problem 2.2

One mole of gas in a closed system undergoes a four-step thermodynamic cycle. Use the data given in the following table to determine numerical values for the missing quantities, i.e. fill in the blanks.

Step	$\Delta U^s \text{ (J)}$	$Q \text{ (J)}$	$W \text{ (J)}$
1→2	-100	?	-7000
2→3	?	-2800	?
3→4	?	-700	200
4→1	3500	?	?
1→2→3→4→1	?	?	-1200

Solution:

a) For this system, using the first law of Thermo I, and assuming $\Delta KE = \Delta PE = 0$

$$\Delta U = Q + W$$

Step 1→2:

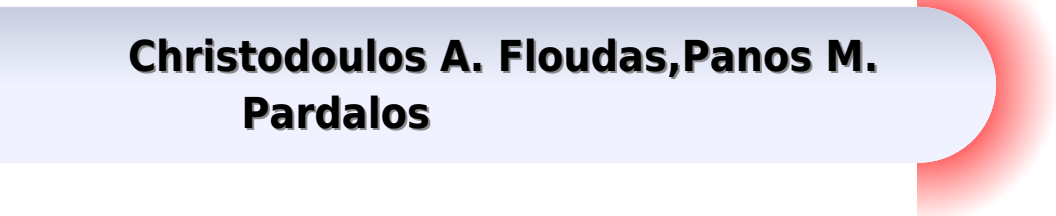
$$\Delta U^{1 \rightarrow 2} = Q + W \rightarrow -100 = Q - 7000 \rightarrow Q^{1 \rightarrow 2} = 6900 \text{ J}$$

Step 3→4:

$$\Delta U^{3 \rightarrow 4} = Q + W = -700 + 200 = -500 \text{ J}$$

Thermodynamics 2paper Solution

**Christodoulos A. Floudas, Panos M.
Pardalos**



Thermodynamics 2paper Solution:

Thermodynamics of Solutions Eli Ruckenstein, Ivan L. Shulgin, 2009-06-17 This book consists of a number of papers regarding the thermodynamics and structure of multicomponent systems that we have published during the last decade. Even though they involve different topics and different systems, they have something in common which can be considered as the signature of the present book. First, these papers are concerned with difficult or very nonideal systems, i.e. systems with very strong interactions, e.g. hydrogen bonding between components or systems with large differences in the partial molar volumes of the components, e.g. the aqueous solutions of proteins or systems that are far from normal conditions, e.g. critical or near critical mixtures. Second, the conventional thermodynamic methods are not sufficient for the accurate treatment of these mixtures. Last but not least, these systems are of interest for the pharmaceutical, biomedical, and related industries. In order to meet the thermodynamic challenges involved in these complex mixtures, we employed a variety of traditional methods but also new methods such as the fluctuation theory of Kirkwood and Buff and ab initio quantum mechanical techniques. The Kirkwood-Buff theory is a rigorous formalism which is free of any of the approximations usually used in the thermodynamic treatment of multicomponent systems. This theory appears to be very fruitful when applied to the above-mentioned difficult systems.

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