Q1. What are the units for each of the following if the concentrations are expressed in moles per liter and the time in seconds?

Rate a second-order reaction	Answer I Question I	Liferoha	
Rate constant for a zero-order reaction	Answer 2 Question 1	Liveroks	-
Rate constant for a first-order reaction	Answer 3 Question 1	1/s	-
Proposition of the Control of the Co	Answer 4 Question 1	mostrs	-

©2. The minimum energy required to transform reactants into an activated complex is called ______.

- a. Kinetic energy
- D. Heat
- c. Activation energy
- d. Potential energy

Q3. A block of copper is heated to 441 °C and then is dropped into a coffee cup calorimeter containing 292 g of water at 25.0 °C. The final temperature of the system is 45.1 °C. What is the mass of the copper block to the nearest gram (1 decimal place)? (Assume all heat is transferred to the water). Write your answer in the box, and do NOT include units.

Specific heat capacity of copper = 0.383 J/ $(g^{*\circ}C)$

Specific heat capacity of water = $4.18 \text{ J/ (g}^{*\circ}\text{C)}$

Answer: 161.8

 $_{\odot}4.2.00M$ CO and 2.00M H₂O are mixed in a sealed container and the system is left to reach equilibrium. Use the following information to calculate the equilibrium constant, Kc if at equilibrium the concentration of CO₂ is 0.73M

$$CO(g) + H_2O(g) \Rightarrow CO_2(g) + H_2(g)$$

Give your answer in decimal form to 2 significant figures e.g. 0.45

Answer: 6.33

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- 1. Understanding the eBook Thermodynamics Test Chemistry
 - The Rise of Digital Reading Thermodynamics Test Chemistry
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Thermodynamics Test Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Test Chemistry
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Thermodynamics Test Chemistry
 - Personalized Recommendations
 - Thermodynamics Test Chemistry User Reviews and Ratings
 - Thermodynamics Test Chemistry and Bestseller Lists
- 5. Accessing Thermodynamics Test Chemistry Free and Paid eBooks

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- Thermodynamics Test Chemistry eBook Subscription Services
- Thermodynamics Test Chemistry Budget-Friendly Options
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 - Thermodynamics Test Chemistry Compatibility with Devices
 - Thermodynamics Test Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Thermodynamics Test Chemistry
 - Interactive Elements Thermodynamics Test Chemistry
- 8. Staying Engaged with Thermodynamics Test Chemistry
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Test Chemistry
- 9. Balancing eBooks and Physical Books Thermodynamics Test Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Test Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermodynamics Test Chemistry
 - \circ Setting Reading Goals Thermodynamics Test Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermodynamics Test Chemistry
 - Fact-Checking eBook Content of Thermodynamics Test Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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