

Chemical Engineering Thermodynamics

Ch. 20 Exam || Questions & Answers

(100% Correct)

thermochemical equation - ANSWER - A chemical equation that shows the heat involved for the amounts of substances specified

spontaneous change - ANSWER - A change that occurs by itself, without an ongoing input of external energy

entropy (S) - ANSWER - A thermodynamic quantity related to the # of ways the energy of a system can be dispersed through the motions of its particles

second law of thermodynamics - ANSWER - A process occurs spontaneously in the direction that increases entropy of the universe

third law of thermodynamics - ANSWER - Entropy of a perfect crystal is zero at 0 K

standard molar entropy (S°) - ANSWER - The entropy of 1 mol of a substance in its standard state

standard entropy of reaction ($\Delta S^\circ_{\text{rxn}}$) - ANSWER - The entropy change that occurs when all components are in their standard states

free energy (G) - ANSWER - A thermodynamic quantity that is the difference between the enthalpy and the product of the absolute temperature and the entropy: $G = H - TES$

standard free energy change (ΔG°) - ANSWER - The free energy change that occurs when all components are in their standard states

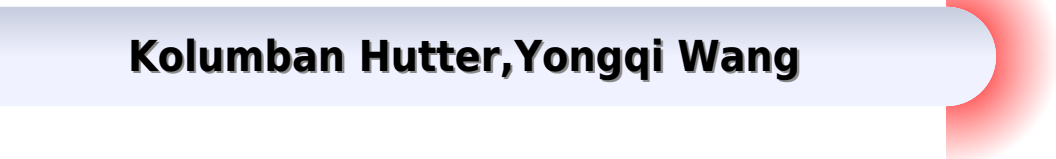
standard free energy of formation (ΔG°_f) - ANSWER - The standard free energy change that occurs when 1 mol of a compound is made from its elements with all components in their standard states

coupling of reactions - ANSWER - The pairing of reactions of which one releases enough free energy for the other to occur

adenosine triphosphate (ATP) - ANSWER - A high-energy molecule that serves most commonly as a store and source of energy in organisms

Thermodynamics Ch 27 Answers

Kolumban Hutter, Yongqi Wang



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