

# Thermochemistry Problems

## The First Law

1. The complete combustion of acetic acid,  $\text{HC}_2\text{H}_3\text{O}_2$ , to form water,  $\text{H}_2\text{O}$ , and  $\text{CO}_2$ , at constant pressure releases 871.7 kJ of heat per mol of acetic acid. Write a balanced thermochemical equation for this reaction. Draw an enthalpy diagram for the reaction.
2. Consider the following reaction, which occurs at standard state:  
$$2\text{Cl}_{(g)} \rightarrow \text{Cl}_{2(g)} \quad \Delta H^\circ = -243.4 \text{ kJ}$$
  - a. Which has the higher, more positive, enthalpy under these conditions,  $2\text{Cl}_{(g)}$  or  $\text{Cl}_{2(g)}$ ?
  - b. Consider the following reaction:  $2 \text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2 \text{MgO}_{(s)}$   $\Delta H = -1204 \text{ kJ}$ .
  - c. Is the reaction exothermic or endothermic?
  - d. Calculate the amount of heat transferred when 2.4 g of  $\text{Mg}_{(s)}$  reacts at constant pressure. How many grams of  $\text{MgO}$  are produced during an enthalpy change of 96.0 kJ?
  - e. How many kilo joules of heat are absorbed when 7.50 g of  $\text{MgO}_{(s)}$  are decomposed into  $\text{Mg}_{(s)}$  and  $\text{O}_{2(g)}$  at constant pressure?

## Enthalpy

## Calorimetry Problems

3. Determine the specific heat of a sample of Cu from the fact that 64.0J are needed to raise the temperature of 15.0 g of Cu metal from 22.0°C to 33.0°C. [0.388 J/g°C]
4. The specific heat of a sample of copper is 0.385 J/g-K. How many joules of heat are necessary to raise the temperature of a 1.42 kg block of copper from 25.0 °C to 88.5°C?
5. A 50.0-g sample of water at 100.00°C was placed in an insulated cup. Then 25.3-g of zinc at 25.00°C was added to the water. The temperature of the water dropped to 96.68°C. What is the specific heat of the zinc?[0.388 J/g°C]
6. When a 6.50 g sample of solid NaOH dissolves in 100.0g of water in a coffee-cup calorimeter, the temperature rises from 21.6°C to 37.8°C. Calculate the  $\Delta H$  (in kJ/mol NaOH) for the solution process. (Assume the specific heat of the solution formed is 4.18 J/g-K)  
$$\text{NaOH}_{(s)} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$$
7. A house is designed to have passive solar energy features, Brickwork is to be incorporated into the interior of the house to act as a heat absorber. Each brick

# Thermochemistry Problems And Solutions

**N Colangelo**



## **Thermochemistry Problems And Solutions:**

## Unveiling the Energy of Verbal Artistry: An Emotional Sojourn through **Thermochemistry Problems And Solutions**

In some sort of inundated with screens and the cacophony of immediate connection, the profound power and emotional resonance of verbal art frequently diminish into obscurity, eclipsed by the continuous assault of noise and distractions. However, located within the musical pages of **Thermochemistry Problems And Solutions**, a charming work of fictional beauty that impulses with natural thoughts, lies an remarkable journey waiting to be embarked upon. Published by a virtuoso wordsmith, that interesting opus guides visitors on a mental odyssey, softly exposing the latent potential and profound influence embedded within the delicate internet of language. Within the heart-wrenching expanse with this evocative evaluation, we will embark upon an introspective exploration of the book is main themes, dissect its charming writing fashion, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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