

Thermochemistry - Energy Flow and Chemical Changes

1. Determine the amount of heat (in kJ) given off when 1.26×10^5 g of NO_2 are produced according to the equation:



$$\Delta H = -114.6 \text{ kJ/mol}$$

2. Consider the reaction:



$$\Delta H = -184.6 \text{ kJ/mol}$$

If 3 moles of H_2 react with 3 moles of Cl_2 to form HCl , calculate the work done (in joules) against a pressure of 1.0 atm at 25°C . What is ΔE for this reaction? Assume the reaction goes to completion.

3. A quantity of 200. mL of 0.862 M HCl is mixed with 200. mL of 0.431 M $\text{Ba}(\text{OH})_2$ in a constant-pressure calorimeter of negligible heat capacity. The initial temperature of the HCl and $\text{Ba}(\text{OH})_2$ solutions is the same at 20.48°C . For the process:



$$\Delta H = -56.2 \text{ kJ/mol}$$

What is the final temperature of the mixed solution? (Assume that the combined solution has a mass of 400g and a specific heat capacity of $4.184 \text{ J/}^\circ\text{C}\cdot\text{g}$)

4. Given the following data:



Calculate the enthalpy change for the reaction:

5. The standard enthalpy of combustion of ethane gas, $\text{C}_2\text{H}_6(g)$, is -1560 kJ/mol at 298 K . Given the following enthalpies of formation, calculate $\Delta H^\circ_{\text{f}}(\text{C}_2\text{H}_6(g))$.



Thermochemistry Heat And Chemical Changes Answers

Lingjun Ying



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